

5. Predicting Products Worksheet

I. Recognizing Reactions

Identify each reaction below by type (combination, decomposition, combustion, single replacement or double replacement)

1. $2 \text{Al}_2\text{O}_3 (\text{s}) \rightarrow 4 \text{Al} (\text{s}) + 3 \text{O}_2 (\text{g})$ _____
2. $\text{S}_8 (\text{s}) + 16 \text{Cl}_2 (\text{g}) \rightarrow 8 \text{SCl}_2 (\text{g})$ _____
3. $\text{CH}_4 (\text{g}) + 2 \text{O}_2 (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + 2 \text{H}_2\text{O} (\text{g})$ _____
4. $\text{NH}_4\text{NO}_3 (\text{s}) \rightarrow \text{N}_2\text{O} (\text{g}) + 2 \text{H}_2\text{O} (\text{l})$ _____
5. $\text{Fe} (\text{s}) + \text{Na}_3\text{PO}_4 (\text{aq}) \rightarrow \text{FePO}_4 (\text{aq}) + 3 \text{Na} (\text{s})$ _____
6. $\text{Ca}(\text{NO}_3)_2 (\text{aq}) + \text{Ag}_2\text{SO}_4 (\text{aq}) \rightarrow 2 \text{AgNO}_3 (\text{aq}) + \text{CaSO}_4 (\text{s})$ _____

II. Single Replacement Reactions

Write the products of the following single replacement reactions and balance the equations:

1. $\text{Li} (\text{s}) + \text{NaCl} (\text{aq}) \rightarrow$
2. $\text{K} (\text{s}) + \text{CaF}_2 (\text{aq}) \rightarrow$
3. $\text{Al} (\text{s}) + \text{MnO} (\text{aq}) \rightarrow$
4. $\text{Sr} (\text{s}) + \text{Cu}(\text{OH})_2 (\text{aq}) \rightarrow$
5. $\text{Ag} (\text{s}) + \text{Au}(\text{NO}_3)_3 (\text{aq}) \rightarrow$

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L Darling-Hammond



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