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substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to

produce water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ hydrochloric

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equilibrium result that can always be formulated as follows $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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species that is capable of accepting a proton which requires a lone pair of electrons to bond to the H^+ water is amphoteric

which means it can act as both a brønsted lowry acid and a brønsted lowry base

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accepts a proton back again it obviously reforms the acid HA

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