

CHEMISTRY AND PHYSICS PRACTICALS

CHEMISTRY

- You are provided with ammonium chloride (NH_4Cl), sodium hydroxide pellets (NaOH), measuring cylinder, thermometer, stirring rod, distilled water, spatula.
 - Pour 50 cm^3 of distilled water into one test tube.
 - Measure the initial temperature in the test tube and record in the table of results.
 - Add half a spatula of ammonium chloride (NH_4Cl) into the test tube and stir gently.
 - Measure the final temperature of the solution in the test tube and record the result.
 - Remove the thermometer from the test tube and rinse it with water.
 - Repeat the steps (a) to (d) with sodium hydroxide pellets and record the results.

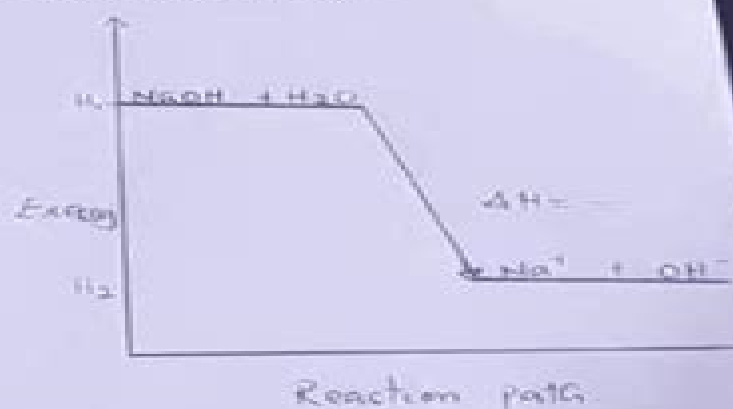
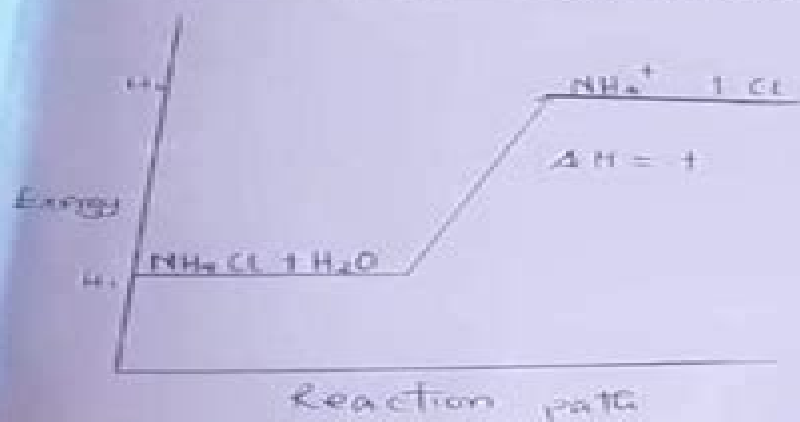
Solution	Initial ($^{\circ}\text{C}$) temperature	Final ($^{\circ}\text{C}$) temperature	Temperature change (final - initial)
NH_4Cl	28°C	25°C	-3°C
NaOH	28°C	45°C	17°C

- Classify the reactions as exothermic or endothermic.

Water + ammonium chloride Endothermic

Water + sodium hydroxide Exothermic

- Draw energy level diagrams to illustrate the above two reactions.



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