

Equilibrium Practice Problems Report

Show your work and write your answers on this page or on separate paper and submit the pictures.

$$R = 0.08206 \text{ Latm/K}\cdot\text{mol} = 8.3144 \text{ J/K}\cdot\text{mol}$$

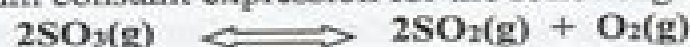
$$K_p = K_c (RT)^{\Delta n_g}$$

1. Write the K_p expression for the reaction



2. Explain what it means when a reaction is at equilibrium. (3 pts.)

3. Write an equilibrium constant expression for the following reaction: (2pts.)



4. Consider the reaction



At equilibrium there is 0.150 mol A_2 , 0.105 mol B_2 and 0.000320 mol AB in a 20.0 L flask at 2300K. Calculate the equilibrium constant, K_c . (4 pts.)

Kp Equilibrium Practice Problems With Answers

WJ Hussar



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