

Chapter 2 Properties of Matter

Section 2.1 Classifying Matter

(pages 38–44)

This section explains how materials are classified as pure substances or mixtures. It discusses types of pure substances and mixtures.

Reading Strategy (page 38)

Summarizing As you read, complete the classification of matter in the diagram below. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

**Pure Substances (page 39)**

- Is the following sentence true or false? Every sample of a pure substance has exactly the same composition and the same properties. true
- What are the two categories of pure substances?
a. Elements b. Compounds

Elements (pages 39–40)

- What is an element? An element is a substance that cannot be broken down into simpler substances.
- Is the following sentence true or false? The smallest particle of an element is an atom. true
- Why does an element have a fixed, uniform composition? An element has a fixed composition because it contains only one type of atom.
- Circle the letter before each element that is a gas at room temperature.
a. carbon b. oxygen
c. mercury d. nitrogen

Match each element to its correct symbol.

Element	Symbol
<u>b.</u> 7. aluminum	a. C
<u>c.</u> 8. gold	b. Al
<u>a.</u> 9. carbon	c. Au

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web purposes objectives in this lab you will be able to make atom models that show stable atoms or ions use given information about subatomic particles to a identify an element and its position on the periodic table b draw models of atoms using build an atom phet sim c determine if the model is for a neutral atom or an ion

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web what particles are located in the center of an atom click the card to flip protons and neutrons click the card to flip 1 22

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web using zoom breakout groups you can work together with your group members to answer the questions in the atomic theory lab report this is an individual assignment so while you can work with your peers all work must be in your own words use the phet simulation below to answer the questions in your lab report

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web the charge of an atom is determined by the ratio of protons to electrons build an atom 1 f 6 play with the simulation to discover what affects the mass number of your atom or ion click on the green sign next to mass to reveal the balance then continue to play the same way you were before

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web try out different models by shooting light at the atom check how the prediction of the model matches the experimental results how did scientists figure out the structure of atoms without looking at them