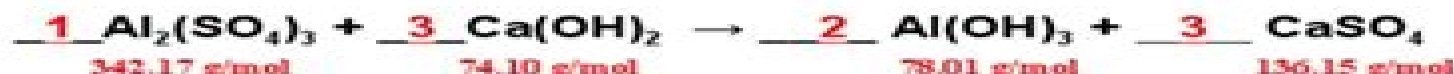


Name: _____ Date: _____ Pd: _____

MIXED Stoichiometry Practice Review Problems

STEP 1: The following reaction will be used for this entire worksheet, please complete the balancing before going on.



STEP 2: Calculator the molar mass of each compound in the balanced reaction and write the values above the balanced reaction.

STEP 3: For aluminum sulfate, find the % of the metal in the compound. Please be sure to show your work. (Hint: WS.1 work)

Al is the only metal, S and O are nonmetals. There are no metalloids in the formula.

$$\% \text{ Al} = \frac{\text{total mass Al}}{\text{molar mass Al}_2(\text{SO}_4)_3} \times 100 \% = \frac{53.96 \text{ g}}{342.17 \text{ g}} \times 100 \% = 15.77 \%$$

STEP 4: Complete the following problems using **a factor chart with units & chemicals**. These problems could require ONE ratio, TWO ratios or even THREE ratios. Look for clues & use your mole map if necessary.

PROBLEMS

1) How many **grams of calcium sulfate** are produced from reacting **2.59 moles of aluminum sulfate**?

$$2.59 \text{ mol Al}_2(\text{SO}_4)_3 \times \frac{3 \text{ mol CaSO}_4}{1 \text{ mol Al}_2(\text{SO}_4)_3} \times \frac{136.15 \text{ g CaSO}_4}{1 \text{ mol CaSO}_4} = 1057.89 \text{ g CaSO}_4$$

moles A \longrightarrow moles B \longrightarrow grams B

Mixed Problems Stoichiometry With Answers

S Marginson



Mixed Problems Stoichiometry With Answers:

The book delves into Mixed Problems Stoichiometry With Answers. Mixed Problems Stoichiometry With Answers is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Mixed Problems Stoichiometry With Answers, encompassing both the fundamentals and more intricate discussions.

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Table of Contents Mixed Problems Stoichiometry With Answers

1. Understanding the eBook Mixed Problems Stoichiometry With Answers
 - The Rise of Digital Reading Mixed Problems Stoichiometry With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mixed Problems Stoichiometry With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mixed Problems Stoichiometry With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mixed Problems Stoichiometry With Answers
 - Personalized Recommendations
 - Mixed Problems Stoichiometry With Answers User Reviews and Ratings
 - Mixed Problems Stoichiometry With Answers and Bestseller Lists
5. Accessing Mixed Problems Stoichiometry With Answers Free and Paid eBooks
 - Mixed Problems Stoichiometry With Answers Public Domain eBooks
 - Mixed Problems Stoichiometry With Answers eBook Subscription Services
 - Mixed Problems Stoichiometry With Answers Budget-Friendly Options
6. Navigating Mixed Problems Stoichiometry With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Mixed Problems Stoichiometry With Answers Compatibility with Devices
 - Mixed Problems Stoichiometry With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mixed Problems Stoichiometry With Answers
 - Highlighting and Note-Taking Mixed Problems Stoichiometry With Answers
 - Interactive Elements Mixed Problems Stoichiometry With Answers

8. Staying Engaged with Mixed Problems Stoichiometry With Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mixed Problems Stoichiometry With Answers
9. Balancing eBooks and Physical Books Mixed Problems Stoichiometry With Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mixed Problems Stoichiometry With Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Mixed Problems Stoichiometry With Answers
 - Setting Reading Goals Mixed Problems Stoichiometry With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Mixed Problems Stoichiometry With Answers
 - Fact-Checking eBook Content of Mixed Problems Stoichiometry With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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