

2. Sodium thiosulfate solution, $\text{Na}_2\text{S}_2\text{O}_3(\text{aq})$, and hydrochloric acid, $\text{HCl}(\text{aq})$, react to produce solid sulfur as in the equation below.



The following results to determine the initial rate were obtained:

Experiment	$[\text{S}_2\text{O}_3^{2-}(\text{aq})] / \text{mol dm}^{-3}$	$[\text{H}^+(\text{aq})] / \text{mol dm}^{-3}$	Initial rate / $\text{mol dm}^{-3} \text{s}^{-1}$
1	0.200	2.00	0.036
2	0.200	1.00	0.036
3	0.100	1.00	0.018

- (a) Deduce, with a reason, the order of reaction with respect to each reactant.

[2]

- (b) State the rate expression for this reaction.

[1]

- (c) Determine the value of the rate constant, k , and state its units.

[2]

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DJ Losen



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