

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<hr/> 58.5 g

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<hr/> 84.0

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<hr/> 159.6

# Modeling Chemistry Unit 5 Test B V2 0

**R R Bowker Publishing, Bowker**



**Modeling Chemistry Unit 5 Test B V2 0:**

This book delves into Modeling Chemistry Unit 5 Test B V2 0. Modeling Chemistry Unit 5 Test B V2 0 is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Modeling Chemistry Unit 5 Test B V2 0, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Modeling Chemistry Unit 5 Test B V2 0
- Chapter 2: Essential Elements of Modeling Chemistry Unit 5 Test B V2 0
- Chapter 3: Modeling Chemistry Unit 5 Test B V2 0 in Everyday Life
- Chapter 4: Modeling Chemistry Unit 5 Test B V2 0 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Modeling Chemistry Unit 5 Test B V2 0. The first chapter will explore what Modeling Chemistry Unit 5 Test B V2 0 is, why Modeling Chemistry Unit 5 Test B V2 0 is vital, and how to effectively learn about Modeling Chemistry Unit 5 Test B V2 0.

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4. In chapter 3, this book will examine the practical applications of Modeling Chemistry Unit 5 Test B V2 0 in daily life. This chapter will showcase real-world examples of how Modeling Chemistry Unit 5 Test B V2 0 can be effectively utilized in everyday scenarios.

5. In chapter 4, the author will scrutinize the relevance of Modeling Chemistry Unit 5 Test B V2 0 in specific contexts. This chapter will explore how Modeling Chemistry Unit 5 Test B V2 0 is applied in specialized fields, such as education, business, and technology.

6. In chapter 5, this book will draw a conclusion about Modeling Chemistry Unit 5 Test B V2 0. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Modeling Chemistry Unit 5 Test B V2 0.

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