

# How To Find Percent Yield

$$\% \text{ yield} = \frac{\text{actual}}{\text{theoretical}} \times 100\%$$

Actual: given in the question

Theoretical : calculated by you

Transformation  
TUTORING

Example: When 10 grams of Zn reacted with 20 grams of HCl, 0.250g of hydrogen gas were obtained in the lab. What is the % yield?



$\text{Zn:}$  10g Zn  $\times$   $\frac{1 \text{ mol Zn}}{65.38 \text{ g Zn}}$   $\times$   $\frac{1 \text{ mol H}_2}{1 \text{ mol Zn}}$   $= 0.153 \text{ mol H}_2$

limiting SMALLER

$\text{HCl:}$  20g HCl  $\times$   $\frac{1 \text{ mol HCl}}{36.458 \text{ g HCl}}$   $\times$   $\frac{1 \text{ mol H}_2}{2 \text{ mol HCl}}$   $= 0.274 \text{ mol H}_2$

$0.153 \text{ mol H}_2 \times \frac{2.016 \text{ g}}{1 \text{ mol H}_2} = 0.308 \text{ g H}_2$  theoretical yield of hydrogen gas

$$\% \text{ yield} = \frac{\text{actual}}{\text{theoretical}} \times 100\%$$

$$\% \text{ yield} = \frac{0.250 \text{ g}}{0.308 \text{ g}} \times 100\% = 81.1\%$$

# Limiting Reactant And Percent Yield Test

**Steve Owen, Chris Martin**



## Limiting Reactant And Percent Yield Test:

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