CECIL CONVENT SCHOOL

PERIODIC ASSESSMENT-2

Class 11 - Chemistry

Time Allowed: 1 hour

General Instructions:

Maximum Marks: 20

TAIL

111

121

121

131

- 1. There are 11 questions in this question paper.
- SECTION A Q. No. 1 to 5 are very short answer questions carrying 1 mark each.
- 3. SECTION B Q. No. 6 to 8 are short answer questions carrying 2 marks each.
- 4. SECTION C- Q. No. 9 to 11 is case based question carrying 3 marks each.
- 5. All questions are compulsory.
- 6. Attempt questions in order.
- 1. Water can be lifted into the water tank at the top of the house with the help of a pump. Then why is it not considered to be spontaneous?
- Neither q nor W is a state function but q + W is a state function. Explain why?
- 3. What is the enthalpy change for an adiabatic process? [1]
- 4. State the law of chemical equilibrium. . [1]
- 5. A tank is full of water. Water is coming in as well as going out at same rate. What will happen to level of water in a tank? What is name given to such state?
- 6. Calculate the standard enthalpy of formation of CH₃OH(I) from the following data:
 - i. $CH_3OH(1) + \frac{3}{2}O_2(g) \longrightarrow CO_2(g) + 2H_2O(1)$; $\Delta_r H^{\odot} = -726 \text{ kJmol}^{-1}$
 - ii. $C_{(graphite)} + O_2(g) \longrightarrow CO_2(g)$; $\Delta_e H^{\odot} = -393 \text{ k/mol}^{-1}$
 - iii. $H_2(g) + \frac{1}{2}O_2(g) \longrightarrow H_2O(1); \Delta_f H^\odot = -286 \text{ kJmol}^{-1}$
- The equilibrium constant for a reaction is 10. What will be the value of ΔGΘ? R = 8.314JK⁻¹mol⁻¹, T = 300 K.
- 8. Write the equilibrium constant (K_c) expression for the following reactions.
 - (i) $Cu^{2+}(aq) + 2Ag(s) = Cu(s) + 2Ag^{+}(aq)$
 - (ii) $4HCl(g) + O_2(g) = 2Cl_2(g) + 2H_2O(g)$
- 9. i. Why standard entropy of an elementary substance is not zero whereas standard enthalpy of formation is taken as zero?
 - ii. Under what conditions will the reaction occur, if
 - a, both ΔH and ΔS are positive
 - b. both ΔH and ΔS are negative
- 10. Dihydrogen gas is obtained from the natural gas by partial oxidation with steam as per following endothermic reaction:

$$CH_4(g) + H_2O(g) \rightleftharpoons CO(g) + 3H_2(g)$$

a. Write the expression for Kp for the above reaction.

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R Sandford

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