

- Carry out as many accurate titrations as you think necessary to obtain consistent results.
- Make sure any recorded results show the precision of your practical work.
- Record in a suitable form below all of your burette readings and the volume of **FA 3** added in each accurate titration.

I	
II	
III	
IV	
V	
VI	
VII	
VIII	

[8]

- (b) From your accurate titration results, obtain a suitable value for the volume of **FA 3** to be used in your calculations.
Show clearly how you obtained this value.

Volume of **FA 3** required = cm³ (1)

2. Calculations

- (a) Calculate the number of moles of sodium carbonate present in the volume of **FA 3** used in your titration.

moles of $\text{Na}_2\text{CO}_3 = \dots\dots\dots$ mol (2)

- (b) Write the equation for the reaction of **FA 3** (Na_2CO_3) with **FA 2** (HCl).

..... (3)

Use your answer to (b) to deduce the number of moles of **FA 2** present in the volume you calculated in (b).

moles of **FA 2** = mol (2)

June 2013 Chemistry 9701 Paper 35 Question

Jessica J Manson



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