

MASTER FORMULA SHEET

CHAPTER 10:

PV = nRT (IDEAL GAS)

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \text{ (COMBINED GAS LAW)}$$

$$J = \frac{C_M}{RT} \text{ (CALCULATING DEBYE OR MOLAL MASS)}$$

$$P_1 = P_2 = P_3 = \dots \text{ (TOTAL PRESSURE OR PARTIAL PRESSURES)}$$

$$P_1 = \left(\frac{n_1}{n_2}\right) P_2 = X_1 P_2 \text{ (RELATING T. PART. TO MOLE FRACTION)}$$

$$M_{rms} = \sqrt{\frac{3RT}{M}} \text{ (DEFINITION OF R.M.S. VELOCITY)}$$

$$\frac{r_1}{r_2} = \sqrt{\frac{M_2}{M_1}} \text{ (RATE OF EFFUSION TO MOLE WEIGHTS)}$$

CHAPTER 11: IN INTERMOLECULAR FORCES

LPE < DIPOLE-DIPOLE < H-BONDING < VAN DER WAALS

IF WEIGHT OF A MOLECULE IS MUCH > ANOTHER, DISPERSION RULES

$$\ln P = \frac{-\Delta H_{vap}}{RT} + C \text{ (CLAUDETT-CLAUDETT)}$$

WITH GRAPH: ΔH_{vap} C-SLOPE = R

CHAPTER 12:

REC. CAT. / FORM. UNIT = AN. COORD. N

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$$\frac{4\pi r^3}{3} \text{ (VOL. OF SPHERE)}$$

PACK. CUBE, BODY-CENT. CUBE, FACE-CENT. CUBE, LATTICE

LIATUM / 2 ATOMS, 4 ATOMS, 8 ATOMS

CHAPTER 15:

$$K_c = \frac{[D]^d [E]^e}{[A]^a [B]^b} \text{ (EQUILIBRIUM CONSTANT)}$$

$$K_p = \frac{(P_D)^d (P_E)^e}{(P_A)^a (P_B)^b} \text{ (EQUILIBRIUM CONSTANT)}$$

$$K_p = K_c (RT)^{\Delta n} \text{ (RELATING THE TWO EQS)}$$

$$Q_c = \frac{[D]^d [E]^e}{[A]^a [B]^b} \text{ (REACTION QUOTIENT)}$$

$$Q_c = \frac{[D]^d [E]^e}{[A]^a [B]^b} \text{ (PREDICTING DIRECTION)}$$

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CHAPTER 17:

$$\Delta S = 4R \ln T \text{ (FOR REVERSIBLE PROCESS)}$$

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