

6. When one gram acid of  $\text{KMnO}_4$  reacts with  $\text{HCl}$ , the volume of solution liberated at NTP will be
- (a) 11.2 litres (b) 22.4 litres  
(c) 44.8 litres (d) 66.8 litres
7. When  $\text{MnO}_4^-$  ions reacts with  $\text{Fe}^{2+}$  ions in acid solution the  $\text{Fe}^{3+}$  is liberated. The equivalent weight of  $\text{KMnO}_4$  in this reaction is
- (a) 500 (b) 1500  
(c) 300 (d) 2000
8. 54 g. of hydrogen peroxide is present in 1000 ml. of solution. This solution is called
- (a) 10 vol solution (b) 20 vol solution  
(c) 30 vol solution (d) 50 vol solution
9. The number of moles of  $\text{KMnO}_4$  that will be needed to react completely with one mole of ferrous oxalate in acidic solution is
- (a) 10 (b) 20  
(c) 40 (d) 2
10. To prepare a solution that is 0.05 M  $\text{KCl}$  starting with 100 mL of 0.40 M  $\text{KCl}$
- (a) add 0.30 g  $\text{KCl}$   
(b) add 50 mL of water  
(c) add 0.30 mol  $\text{KCl}$   
(d) evaporate 10 mL water
11. In hot alkaline solution,  $\text{Br}_2$  disproportionates to  $\text{Br}^-$  and  $\text{BrO}_2^-$
- $$3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + \text{BrO}_2^- + 3\text{H}_2\text{O}$$
- Atomic equivalent weight of  $\text{Br}_2$  is (molecular weight = 160)
- (a)  $\frac{16}{5}$  (b)  $\frac{16}{3}$   
(c)  $\frac{160}{5}$  (d)  $\frac{160}{3}$
12.  $(\text{Fe}^{2+})$  in a solution prepared by adding 20.00 mL of 0.10 M  $\text{FeSO}_4$  with 70 mL of 0.10 M  $\text{Na}_2\text{SO}_4$  is
- (a) 0.020 M (b) 0.140 M  
(c) 0.010 M (d) 0.200 M
13. When 80 mL of 0.100 M  $\text{HCl}$  is mixed with 120 mL of 0.10 M  $\text{KOH}$ , the resultant solution is the same as a solution of
- (a) 0.10 M  $\text{HCl}$  and 0.02 M  $\text{HCl}$   
(b) 0.08 M  $\text{HCl}$   
(c) 0.08 M  $\text{HCl}$  and 0.01 M  $\text{KOH}$   
(d) 0.08 M  $\text{HCl}$  and 0.01 M  $\text{HCl}$
14. 50 mL of  $\text{H}_2\text{O}_2$  solution contains strength = x% required 10 mL of  $\text{KMnO}_4$  solution in acidic medium. Hence x is
- (a) 0.20 (b) 0.1  
(c) 0.1 (d) 0.0
15. Molarity of 1.0 M  $\text{H}_2\text{SO}_4$  (d = 1.8 mL/cm<sup>3</sup>) is
- (a) 36 mol kg<sup>-1</sup> (b) 300 mol kg<sup>-1</sup>

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### **Table of Contents Multiple Choice Questions On Inorganic Chemistry**

1. Understanding the eBook Multiple Choice Questions On Inorganic Chemistry
  - The Rise of Digital Reading Multiple Choice Questions On Inorganic Chemistry
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Inorganic Chemistry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions On Inorganic Chemistry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Inorganic Chemistry
  - Personalized Recommendations
  - Multiple Choice Questions On Inorganic Chemistry User Reviews and Ratings
  - Multiple Choice Questions On Inorganic Chemistry and Bestseller Lists

5. Accessing Multiple Choice Questions On Inorganic Chemistry Free and Paid eBooks
  - Multiple Choice Questions On Inorganic Chemistry Public Domain eBooks
  - Multiple Choice Questions On Inorganic Chemistry eBook Subscription Services
  - Multiple Choice Questions On Inorganic Chemistry Budget-Friendly Options
6. Navigating Multiple Choice Questions On Inorganic Chemistry eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions On Inorganic Chemistry Compatibility with Devices
  - Multiple Choice Questions On Inorganic Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Inorganic Chemistry
  - Highlighting and Note-Taking Multiple Choice Questions On Inorganic Chemistry
  - Interactive Elements Multiple Choice Questions On Inorganic Chemistry
8. Staying Engaged with Multiple Choice Questions On Inorganic Chemistry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions On Inorganic Chemistry
9. Balancing eBooks and Physical Books Multiple Choice Questions On Inorganic Chemistry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions On Inorganic Chemistry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions On Inorganic Chemistry
  - Setting Reading Goals Multiple Choice Questions On Inorganic Chemistry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions On Inorganic Chemistry
  - Fact-Checking eBook Content of Multiple Choice Questions On Inorganic Chemistry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning



- Utilizing eBooks for Skill Development
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