ENGAGE

Why does the atomic number but not the mass number change in positron emission?

SAMPLE PROBLEM 5.4 Writing a Nuclear Equation for Positron Emission

TRY IT FIRST

Write the balanced nuclear equation for manganese-49, which decays by emitting a positron.

SOLUTION GUIDE

ANALYZE THE PROBLEM	Given	Need	Connect
	Mn-49, positron	balanced nuclear	mass number, atomic
	emission	equation	number of new nucleus

STEP 1 Write the incomplete nuclear equation.

$$^{49}_{29}Mm \longrightarrow 7 + ^{9}_{47}e$$

STEP 2 Determine the missing mass number. In the equation, the mass number, 49, is equal to the sum of the mass numbers of the new nucleus and the positron.

$$49 = 7 + 0$$

$$49 - 0 = ?$$

$$49 - 0 = 49$$
 (mass number of new nucleus)

STEP 3 Determine the missing atomic number. The atomic number, 25, must equal the sum of the atomic numbers of the new nucleus and the positron.

$$25 = 7 + 1$$

$$25 - 1 = 7$$

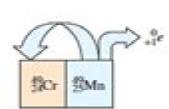
STEP 4 Determine the symbol of the new nucleus. On the periodic table, the element that has atomic number 24 is chromium, Cr. The atomic symbol for this isotope of Cr is written 50r.

SIEP 5 Complete the nuclear equation.

STUDY CHECK S.4.

Write the balanced nuclear equation for xenon-118, which undergoes positron emission.

ANSWER



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Tobias Bleicker

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Table of Contents Mastering Chemistry Exercise Answers

- 1. Understanding the eBook Mastering Chemistry Exercise Answers
 - The Rise of Digital Reading Mastering Chemistry Exercise Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mastering Chemistry Exercise Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mastering Chemistry Exercise Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mastering Chemistry Exercise Answers
 - Personalized Recommendations

- Mastering Chemistry Exercise Answers User Reviews and Ratings
- Mastering Chemistry Exercise Answers and Bestseller Lists
- 5. Accessing Mastering Chemistry Exercise Answers Free and Paid eBooks
 - Mastering Chemistry Exercise Answers Public Domain eBooks
 - Mastering Chemistry Exercise Answers eBook Subscription Services
 - Mastering Chemistry Exercise Answers Budget-Friendly Options
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 - Highlighting and Note-Taking Mastering Chemistry Exercise Answers
 - Interactive Elements Mastering Chemistry Exercise Answers
- 8. Staying Engaged with Mastering Chemistry Exercise Answers
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mastering Chemistry Exercise Answers
- 9. Balancing eBooks and Physical Books Mastering Chemistry Exercise Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mastering Chemistry Exercise Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mastering Chemistry Exercise Answers
 - Setting Reading Goals Mastering Chemistry Exercise Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mastering Chemistry Exercise Answers
 - Fact-Checking eBook Content of Mastering Chemistry Exercise Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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