

12 • The Gas Laws**BOYLE'S LAW**

Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held constant. (If one goes up, the other goes down.) We use the formula:

$$P_1 \times V_1 = P_2 \times V_2$$

Solve the following problems (assuming constant temperature). Assume all number are 3 significant figures.

1. A sample of oxygen gas occupies a volume of 250 mL at 740 torr pressure. What volume will it occupy at 800 torr pressure?

$$P_1 V_1 = P_2 V_2$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(740 \text{ torr})(250 \text{ mL})}{(800 \text{ torr})} = \boxed{231 \text{ mL}}$$

2. A sample of carbon dioxide occupies a volume of 3.50 Liters at 125 kPa pressure. What pressure would the gas exert if the volume was decreased to 2.00 liters?

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{(125 \text{ kPa})(3.50 \text{ L})}{(2.00 \text{ L})} = \boxed{219 \text{ kPa}}$$

3. A 2.00-Liter container of nitrogen had a pressure of 3.20 atm. What volume would be necessary to decrease the pressure to 1.00 atm?

$$P_1 = 3.20 \text{ atm} \quad V_1 = 2.00 \text{ L}$$

$$P_2 = 1.00 \text{ atm}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(3.20 \text{ atm})(2.00 \text{ L})}{(1.00 \text{ atm})} = \boxed{6.40 \text{ L}}$$

4. Ammonia gas occupies a volume of 450 mL as a pressure of 720 mmHg. What volume will it occupy at standard pressure (760 mmHg)?

$$P_1 = 720 \text{ mmHg} \quad V_1 = 450 \text{ mL}$$

$$P_2 = 760 \text{ mmHg}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(720 \text{ mmHg})(450 \text{ mL})}{(760 \text{ mmHg})} = \boxed{426 \text{ mL}}$$

5. A 175 mL sample of neon had its pressure changed from 75.0 kPa to 150 kPa. What is its new volume?

$$V_1$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(75.0 \text{ kPa})(175 \text{ mL})}{150 \text{ kPa}} = \boxed{87.5 \text{ mL}}$$

6. A sample of hydrogen at 1.50 atm had its pressure decreased to 0.50 atm producing a new volume of 750 mL. What was the sample's original volume?

$$V_1 = \frac{P_2 V_2}{P_1} = \frac{(0.50 \text{ atm})(750 \text{ mL})}{(1.50 \text{ atm})} = \boxed{250. \text{ mL}}$$

7. Chlorine gas occupies a volume of 1.20 liters at 720 torr pressure. What volume will it occupy at 1 atm pressure?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(720 \text{ torr})(1.20 \text{ L})}{(1 \text{ atm} \times \frac{760 \text{ torr}}{1 \text{ atm}})} = \boxed{1.14 \text{ L}}$$

8. Fluorine gas exerts a pressure of 900 torr. When the pressure is changed to 1.50 atm, its volume is 250 mL. What was the original volume?

$$V_1 = \frac{P_2 V_2}{P_1} = \frac{(1.50 \text{ atm})(250 \text{ mL})}{(900 \text{ torr} \times \frac{1 \text{ atm}}{760 \text{ torr}})} = \boxed{317 \text{ mL}}$$

J Weston Walch Gas Law Problems Answers

Brian Pressley



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