

Mixed Extra Gas Law Practice Problems (Ideal Gas, Dalton's Law of Partial Pressures, Graham's Law)

1. Dry ice is carbon dioxide in the solid state. 1.28 grams of dry ice is placed in a 5.00 L chamber that is maintained at 38.1°C. What is the pressure in the chamber after all of the dry ice has sublimed?

$$PV = nRT$$

$$\frac{1.28 \text{ g } CO_2}{1} \times \frac{1 \text{ mole } CO_2}{44.009 \text{ g } CO_2} = 0.290849 \text{ moles } CO_2$$

$$P = ?$$

$$V = 5.00 \text{ L (3 significant figures)}$$

$$n = 0.290849 \text{ moles (3 significant figures)}$$

$$R = 0.0821 \text{ L} \cdot \text{atm} / \text{mole} \cdot \text{K (infinite significant figures)}$$

$$T = 38.1^\circ\text{C} = 308.1 \text{ K (4 significant figures)}$$

$$P = \frac{nRT}{V} = \frac{(0.290849 \text{ moles}) \left(0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mole} \cdot \text{K}}\right) (308.1 \text{ K})}{5.00 \text{ L}} = 1.47 \text{ atm}$$

If you used a different R, then the answers are:

$$1120 \text{ torr}$$

$$1120 \text{ mm Hg}$$

$$149 \text{ kPa}$$

2. A sample of chlorine gas is loaded into a 0.25 L bottle at standard temperature of pressure. How many moles of bromine gas are in the container? How many grams?

At STP 1 mole = 22.4 L

Molar Mass of Chlorine (remember, it is a diatomic) = 70.906 g/mole

Use factor label

$$\frac{0.25 \text{ L}}{1} \times \frac{1 \text{ mole } Cl_2}{22.4 \text{ L } Cl_2} = 0.011 \text{ moles } Cl_2$$

$$\frac{0.25 \text{ L}}{1} \times \frac{1 \text{ mole } Cl_2}{22.4 \text{ L } Cl_2} \times \frac{70.906 \text{ g } Cl_2}{1 \text{ mole } Cl_2} = 0.79 \text{ g } Cl_2$$

OR - use the Ideal gas law,  $PV = nRT$

$$P = 1 \text{ atm (infinite significant figures)}$$

$$V = 0.25 \text{ L (2 significant figures)}$$

$$R = 0.0821 \text{ L} \cdot \text{atm} / \text{mole} \cdot \text{K (infinite significant figures)}$$

$$T = 273 \text{ K (infinite significant figures)}$$

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**Judd E. Hollander**



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