

Solve Literal Equations

A literal equation is an equation with more than one variable.

To solve a literal equation means to isolate the indicated variable.

We can use inverse operations to isolate the variable.

1. If necessary, simplify each side of the equation by combining like terms.
2. Isolate the indicated variable -
 - a) Use the opposite operation of addition or subtraction to move any constants or variables to the other side.
 - b) Use the opposite operation of multiplication or division to eliminate the coefficient.

Examples:

Solve for x

$$3x + 4y = 7$$

$$3x = 7 - 4y$$

$$x = \frac{7 - 4y}{3}$$

Solve for F

$$C = \frac{5}{9}(F - 32)$$

$$\frac{9}{5}C = F - 32$$

$$\frac{9}{5}C + 32 = F$$

$$F = \frac{9}{5}C + 32$$

Literal Equations And Formulas Student Guide

Shasha Hu



Literal Equations And Formulas Student Guide:

Maths Handbook & Study Guide Grade 10 Kevin Smith, 2017-11-01 The Maths Handbook Study Guide is a comprehensive reference book and set of notes that covers everything in one book The book is written in a clear simple visual and logical manner The colour coding facilitates explanations definitions formulas recaps of previous work hints and ideas It is easy to read easy to understand and it is easy to apply what has been learnt It works in conjunction with all other Maths books It is a welcome addition to the Handbook and Study Guide series The Maths Handbook and Study Guide demystifies Maths and helps students to reach their potential in this challenging subject The sub title of the book is Maths Made Easy and this is what it aims to do Kevin ensures that his work is up to date at all times and that it is suitable for IEB and National Curriculum students There are exercises in the front of the book and solutions to problems at the back **Worksheets and Study Guide for Kaufmann/Schwitters' Algebra for College Students** Kay Haralson, 2000 Algebra Walter Sherwood, 1998 **Suggested Study Guide, First Year Algebra** , 1954 State Manual of the Courses of Study for the High Schools of Oregon Oregon. Office of Superintendent of Public Instruction, 1927 **Intermediate Algebra for College Students (Student Study Guide)** Kelli Jade Hammer, Allen R. Angel, 2003-05 **Study Guide with Solutions to Selected Odd-numbered Exercises for Gustafson and Frisk's Intermediate Algebra** George Grisham, 1992 **Introductory Algebra** Arnold R. Steffensen, Lee Murphy Johnson, 1994 **College Algebra with Applications Solutions and Study Guide** Mutter, 1998-03-06 *Student Handbook Including Webster's New World Dictionary* Lawrence T. Lorimer, Bryan H. Bunch, 1983 **Student Mathematics Handbook** Gerald L. Bradley, Karl J. Smith, 1995 **Study Guide** Vicky Lymbery, 1990 *Handbook of Mathematics for Engineers and Engineering Students* Joseph Claudel, 1906 **Student Handbook** Gorton Carruth, 1993 A Hudson Group book Includes index A one volume short entry encyclopedia containing information about the characters plots authors and other topics found in classic and contemporary books for young readers

Student's Solutions Manual John R. Martin, Washington, 2000 *Intermediate Algebra for College Students* Allen R. Angel, 2000 Includes additional worked out examples additional drill problems and Practice Tests and their answers Important concepts are emphasized **State Manual of the Courses of Study for the High Schools of Oregon ...** Oregon. Office of Superintendent of Public Instruction, 1931 **Teachers' Guide** Ernst Rudolph Breslich, 1925

Handbook on Formative and Summative Evaluation of Student Learning Benjamin Samuel Bloom, John Thomas Hastings, George F. Madaus, 1971 **Holt Algebra 1 2003** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003

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Stoichiometry helps us determine how much substance is needed or is present in things that can be measured. Reactants and products have mass, molecular weight, and chemical formulas. The stoichiometric coefficient or stoichiometric number is the number of molecules that participate in the reaction.

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using the following equation

$$2\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$

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 $\text{Ce} + \text{Zn} + \text{CuCl}_2 \rightarrow \text{ZnCl}_2 + \text{Cu}$
how many moles of $\text{Ce} + \text{Zn} + \text{CuCl}_2$ will be produced from 23.0 g of $\text{Ce} + \text{Zn}$ assuming $\text{Ce} + \text{CuCl}_2$ is available in excess

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 $\text{N}_2 + 3\text{F}_2 \rightarrow 2\text{NF}_3$
 $2\text{C}_6\text{H}_{10} + 17\text{O}_2 \rightarrow 12\text{CO}_2 + 10\text{H}_2\text{O}$
 $3\text{H}_2\text{O} + 1\text{KBr} + 1\text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr}$
using the following equation
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow 2\text{H}_2\text{O} + \text{Na}_2\text{SO}_4$

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 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr}$
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 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow 4\text{H}_2\text{O} + 2\text{GaBr}_3 + 3\text{Na}_2\text{SO}_3 + 1\text{Ga}_2\text{SO}_3 + 3\text{H}_2\text{O} + 3\text{NaBr}$
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