

## Chemistry Unit 7 Review

### 701 – Inquiry Lab

1. Define...
  - a. Direct relationship: Two variables increase/decrease together
  - b. Indirect relationship: One variable increases, the other decreases
2. Identify whether the listed variables have a direct or indirect relationship to each other.
  - a. Moles and volume: Direct
  - b. Temperature and pressure: Direct
  - c. Pressure and volume: Indirect
  - d. Moles and pressure: Direct
3. Using the Ideal Gas Law equation, list 3 ways you could increase the volume of a balloon.  
Increase the moles, increase the temperature, decrease the pressure (more so, the pressure drops as a result of greater volume)
4. Using the Ideal Gas Law equation, list 3 ways you could decrease the pressure within a syringe.  
Decrease the moles, decrease the temperature, increase the volume
5. If the number of moles in a container is doubled, by how much should the pressure increase (assuming everything else is held constant)?  
The pressure should also double

### 702 – Gas Laws Equation

1. What equation should you use when you are given just one set of variables?  
 $PV = nRT$
2. You have a 4.5 L canister of  $O_2$  that you store in a room that is  $20\text{ }^{\circ}\text{C}$ . The pressure of this tank is 13000 kPa. How many moles of  $O_2$  are in this canister?  
24 moles

# Modeling Chemistry U7 Ws 1 V

**Michael Hoelscher**



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