

Hydrogen ^1H

Oxygen ^{16}O

Oxygen is 16x larger than hydrogen

You shouldn't conclude that chemists were able to determine the molar masses of all the elements using this technique. Measurements of the density of the gaseous phase of many of the elements are difficult, if not impossible. However, we are going to see that chemists could use another tool - the *percent composition of compounds* to determine molar masses.

Relative Mass From Compounds

Many substances combine with oxygen to form a type of compound called an oxide. In the previous unit we saw that such combinations often occur in multiple proportions. John Dalton made the assumption that the lowest ratio was a 1:1 combination of elements. For now, we will make a similar assumption. We may have to re-examine this assumption later.

In Unit 4 you examined % composition data of several compounds. For example, 100 g of the simplest oxide of carbon contains 42.9 g of C and 57.1 g of O. To be able to compare masses of elements to the mass of oxygen, it is important to use the same mass of oxygen in all of the ratios. Using a proportion, one can show that 75.0 g of C combines with 100 g of oxygen.

$$\textcircled{A} \quad \frac{42.9\text{gC}}{57.1\text{gO}} = \frac{\text{?}}{100\text{gO}}$$

$$\textcircled{B} \quad \frac{27.3\text{gC}}{72.7\text{gO}} = \frac{\text{?}}{100\text{gO}}$$

$$\frac{57.1\text{gO}}{42.9\text{gC}} = \frac{1.33 \text{ O}_2}{1 \text{ H}}$$

$$\frac{72.7\text{gO}}{27.3\text{gC}} = 2.66 \text{ O}_2$$

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