## Chemistry Unit 8 Worksheet 4 Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

 Calculate the number of moles of potassium chlorate, KClO<sub>3</sub> (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

Equation:	2 KClO <sub>3</sub> +	≥ 2 KCl +	3 02
Before:	1.2		0
Change:	-1.2	+1.2	+1.8
Afters	.0	1.2	1.8

1.2 moles KCRO, needed

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

Equations	Non-	2 HCl >	MgCl <sub>2</sub> +	882
Before:	0.100	7		0
Change:	-0.100	-0.200 +	0.100 +	0.100
Afters		2	0.100	0.100

0.200 protes HCI needed

0.73 g x 1 mole = .0040 mole 183 gr

 Ethane, C<sub>2</sub>H<sub>6</sub> reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

Equation:	2 C2H6	+ 7 O <sub>2</sub>	→ 4 CO₂	+ 6 H <sub>2</sub> O
Before:	2.20	5.5	0	0.0
Change:	-2.20	~ 7.70	+ 4,40	+ 6,60
After:	0	3.5	4.40	6.60

O2 needed: 7,70 mole x 32.0 g = 246 g O2 required I mode

 Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation: Ni(NO3)2 + 2 NaOH --> Ni(OH)2 + 2 NaNO3

> Ni(NO<sub>3</sub>)<sub>2</sub> + 2 NaOH -> Ni(OH)<sub>2</sub> + 2 NaNO<sub>3</sub> Equation: Before: 0.0040 No. 26 -0.0040- O. GOOSIO **-0.0040** +0.06900

N. S. 0.0080 Afflors 0.0640 NaNO<sub>2</sub> produced: 0.0080 mole x 85.0 g = 0.68 g 1 mode

Change:

# **Modeling Chemistry Unit 7 Answer**

**G** Thomas

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# **Table of Contents Modeling Chemistry Unit 7 Answer**

- 1. Understanding the eBook Modeling Chemistry Unit 7 Answer
  - The Rise of Digital Reading Modeling Chemistry Unit 7 Answer
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Modeling Chemistry Unit 7 Answer
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Modeling Chemistry Unit 7 Answer
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Modeling Chemistry Unit 7 Answer
  - Personalized Recommendations
  - Modeling Chemistry Unit 7 Answer User Reviews and Ratings
  - Modeling Chemistry Unit 7 Answer and Bestseller Lists

- 5. Accessing Modeling Chemistry Unit 7 Answer Free and Paid eBooks
  - o Modeling Chemistry Unit 7 Answer Public Domain eBooks
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  - Modeling Chemistry Unit 7 Answer Budget-Friendly Options
- 6. Navigating Modeling Chemistry Unit 7 Answer eBook Formats
  - o ePub, PDF, MOBI, and More
  - Modeling Chemistry Unit 7 Answer Compatibility with Devices
  - Modeling Chemistry Unit 7 Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 7 Answer
  - Highlighting and Note-Taking Modeling Chemistry Unit 7 Answer
  - Interactive Elements Modeling Chemistry Unit 7 Answer
- 8. Staying Engaged with Modeling Chemistry Unit 7 Answer
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Modeling Chemistry Unit 7 Answer
- 9. Balancing eBooks and Physical Books Modeling Chemistry Unit 7 Answer
  - Benefits of a Digital Library
  - $\circ$  Creating a Diverse Reading Collection Modeling Chemistry Unit 7 Answer
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Modeling Chemistry Unit 7 Answer
  - Setting Reading Goals Modeling Chemistry Unit 7 Answer
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Modeling Chemistry Unit 7 Answer
  - Fact-Checking eBook Content of Modeling Chemistry Unit 7 Answer
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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