

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show the given data and find the mass. The mass is underlined (underlined) given.

1. Chlorine gas reacts with a solution of potassium hydroxide, KOH , to form a solid compound and potassium chloride, KCl , and a diatomic oxygen gas.

Reactants:	Cl_2	KOH	
Products:	KCl	O_2	
Masses:	1.00	1.00	1.00
Mass:	1.00	1.00	1.00

Reaction:
KOH + Cl₂ → KCl + O₂

2. A sample of a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. The mass of the magnesium metal is 1.00 g and the mass of the hydrogen gas is 0.10 g.

Reactants:	Mg	HCl	
Products:	MgCl_2	H_2	
Masses:	1.00	1.00	1.00
Mass:	1.00	1.00	1.00

Reaction:
Mg + HCl → MgCl₂ + H₂

3. Chlorine, Cl_2 , reacts with a solution of potassium hydroxide, KOH , to form a solid compound and potassium chloride, KCl , and a diatomic oxygen gas. The mass of the chlorine gas is 1.00 g and the mass of the potassium chloride is 1.00 g.

Reactants:	Cl_2	KOH	
Products:	KCl	O_2	
Masses:	1.00	1.00	1.00
Mass:	1.00	1.00	1.00

Reaction:
 $\text{Cl}_2 + \text{KOH} \rightarrow \text{KCl} + \text{O}_2$

4. Chlorine gas reacts with a solution of potassium hydroxide, KOH , to form a solid compound and potassium chloride, KCl , and a diatomic oxygen gas. The mass of the chlorine gas is 1.00 g and the mass of the potassium chloride is 1.00 g.

Reactants:	Cl_2	KOH	
Products:	KCl	O_2	
Masses:	1.00	1.00	1.00
Mass:	1.00	1.00	1.00

Reaction:
 $\text{Cl}_2 + \text{KOH} \rightarrow \text{KCl} + \text{O}_2$

Modeling Chemistry Unit 8 Answer

Richard Bailey



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