

CHAPTER 14 REVIEW*Acids and Bases***SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

- 1.**
- Name the following compounds as acids:

| | |
|--------------------------------|----------------------------|
| _____ sulfuric acid _____ | a. H_2SO_4 |
| _____ sulfurous acid _____ | b. H_2SO_3 |
| _____ hydrosulfuric acid _____ | c. H_2S |
| _____ perchloric acid _____ | d. HClO_4 |
| _____ hydrocyanic acid _____ | e. hydrogen cyanide |

- 2.**
- _____
- H_2S
- Which (if any) of the acids mentioned in item
- 1**
- are binary acids?

- 3.**
- Write formulas for the following acids:

| | |
|--------------------------------------|----------------------|
| _____ HNO_2 _____ | a. nitrous acid |
| _____ HBr _____ | b. hydrobromic acid |
| _____ H_3PO_4 _____ | c. phosphoric acid |
| _____ CH_3COOH _____ | d. acetic acid |
| _____ HClO _____ | e. hypochlorous acid |

- 4.**
- Calcium selenate has the formula
- CaSeO_4
- .

| | |
|--------------------------------------|---|
| _____ H_2SeO_4 _____ | a. What is the formula for selenic acid? |
| _____ H_2SeO_3 _____ | b. What is the formula for selenous acid? |

- 5.**
- Use an activity series to identify two metals that will not generate hydrogen gas when treated with an acid.

Choose from Cu, Ag, Au, Pt, Pd, or Hg.

- 6.**
- Write balanced chemical equations for the following reactions of acids and bases:

- a. aluminum metal with dilute nitric acid



- b. calcium hydroxide solution with acetic acid



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