

## STOICHIOMETRY PRACTICE WORKSHEET

1. Using the following equation:



Calculate how many grams of iodine are needed to prepare 28.6 grams of ICl by this reaction.

2. Using the following equation:



How many moles and grams of  $\text{KMnO}_4$  are needed for this reaction on 11.4 grams of  $\text{KNO}_2$ ?

3. Using the following equation:



How many moles and grams of oxygen ( $\text{O}_2$ ) are needed to react with 56.8 grams of ammonia by this reaction?

4. Using the following equation:



Calculate the number of moles and the number of grams of iodine ( $\text{I}_2$ ) that can be made this way from 16.4 grams of  $\text{NaIO}_3$ .

# Mixed Stoichiometry Practice

**Victor M. Corman**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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