

## Unit 5 - Ionic Bonding Study Guide

Name: KEY Period:      Date:     

### Review Ionic Bonds

- What two types of elements will transfer electrons to form an ionic bond?  
*Metal & nonmetal*
- What is an ionic bond?  
*An electrostatic force that holds together a cation and anion*
- How can you describe the electrical charge of an ionic compound?  
*Neutral*
- Metals will *lose* all of their valence electrons to form a positively charged *cation*. The oxidation number will be equal to the number of valence electrons.
- Nonmetals will *gain* valence electrons to form a negatively charged *anion*. The oxidation number will be equal to 8 – the # of valence electrons.
- Which elements are least likely to undergo bonding? *Noble Gases*
- Write the oxidation number for the following elements.

a. Cesium	<u>+1</u>	e. Magnesium	<u>+2</u>
b. Potassium	<u>+1</u>	f. Barium	<u>+2</u>
c. Aluminum	<u>+3</u>	g. Nitrogen	<u>-3</u>
d. Iron (III)	<u>+3</u>	h. Iodine	<u>-1</u>
		i. Oxygen	<u>-2</u>
		j. Selenium	<u>-2</u>
- Which alkali metal has the highest ionization energy? *Hydrogen*
- Which alkaline earth metal has the lowest ionization energy? *Radium*
- What properties characterize ionic compounds?
  - High melting point & boiling point*
  - Solid at room temperature*
  - Brittle*
  - Ionic solids can't conduct electricity*
  - Molten ionic compounds can conduct electricity*

### Transferring of Electrons

11. Please draw the transfer of electrons for the following:

a. NaBr



# Ionic Bonding Study Guide

**David E. Sadava, Gordon H.  
Orians, Craig Heller, William K. Purves**



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