

How to Name Chemical Compounds

is it ionic (metal, or does it contain a polyatomic ion that provides an H⁺)

otherwise...

Ionic compound

[cation's name] [anion's name]

main element	<ul style="list-style-type: none"> take element name! ex. Al³⁺: aluminum if transition metal*, include charge in Roman numeral ex. Cr³⁺: chromium (III) 	<ul style="list-style-type: none"> get element's name, chop end, add "-ide" ex. F⁻: fluorine → fluoride⁻
poly-atomic	<ul style="list-style-type: none"> just use its name! ex. NH₄⁺: ammonium 	<ul style="list-style-type: none"> just use its name! ex. CO₃²⁻: carbonate

* exceptions (no Roman numerals):

silver ion: always Ag⁺

zinc ion: always Zn²⁺

cadmium ion: always Cd²⁺

Molecular compound

is it HClO_x?

Acid!
root and numbers

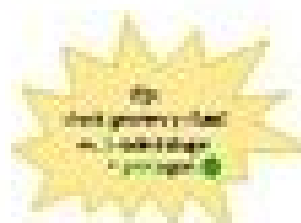
[#]-[1st element name] [#]-[2nd element root]-"ide"

↑ skip if "mono"

↑ skip if 1st element is H

Numerical prefixes

- 1: mono-
- 2: di-
- 3: tri-
- 4: tetra-
- 5: penta-
- 6: hexa-
- 7: hepta-
- 8: octa-
- 9: nona-
- 10: deca-



Acids

Ignore all H⁺, and look at the name of the remaining anion.

Anion	Xide	Xate	Xite
Acid	hydro-X-ic acid	X-ic acid	X-ous acid
	ex. Br ⁻ bromide HBr hydrobromic acid	ex. ClO ₃ ⁻ chlorate HClO ₃ chloric acid	ex. ClO ₂ ⁻ chlorite HClO ₂ chlorous acid
	ex. CN ⁻ cyanide HCN hydrocyanic acid	ex. PO ₄ ³⁻ phosphate H ₃ PO ₄ phosphoric acid	ex. SO ₃ ²⁻ sulfite H ₂ SO ₃ sulfurous acid

Ionic Compounds Study Guide

Susan McMurry



Ionic Compounds Study Guide:

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