

Ionic Equations

Ionic equations show us which **species** (atoms, ions and molecules) are involved in a reaction.

An ionic substance splits up into its ions.

E.g. NaCl forms $\text{Na}^+ + \text{Cl}^-$

This also happens if an acid or soluble base dissolves in water.

Must be (aq) – not (s) or (g)

E.g. HCl (aq) forms $\text{H}^+ + \text{Cl}^-$ $\text{Mg(OH)}_2 \text{ (aq)}$ forms $\text{Mg}^{2+} + 2\text{OH}^-$ but $\text{Ca(OH)}_2 \text{ (s)}$ does not dissociate

Pure elements and covalent substances (except acids) will not form ions.

E.g. Mg , C , CO_2 , H_2

1. Circle the substances that will NOT form ions:

NaCl CO_2 CuCO_3 H_2 SiO_2 Al C_2H_6 Mg(OH)_2 LiCl PbI_2 H_2O

2. Identify the species in these substances:

a) MgSO_4 $\text{Mg}^{2+} + \text{SO}_4^{2-}$

b) CaCl_2

c) Li_2O

d) HNO_3

e) $2\text{Na}_2\text{O}$

f) 2CaCO_3

g) H_2SO_4 (sulfuric acid)

h) HNO_3 (nitric acid)

i) Al_2O_3

j) MgCO_3

k) Fe(OH)_3

l) Na_2SO_4

3. Identify the species in these substances:

a) $\text{Mg} + 2\text{HCl}$ $\text{Mg} + 2\text{H}^+ + 2\text{Cl}^-$

b) $\text{MgCl}_2 + \text{H}_2$

c) $\text{Li} + \text{O}_2$

d) $\text{CaCO}_3 + 2\text{HCl}$

e) $2\text{Na}_2\text{O} + 2\text{Li}$

f) $2\text{Zn} + \text{H}_2\text{SO}_4$

g) $\text{MgO} + \text{CO}_2$

h) $2\text{NaI} + \text{Br}_2$

i) $\text{Ca} + 2\text{HNO}_3$

4. Identify the species in the reactants and products:

a) $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
 $\text{Mg} + 2\text{H}^+ + 2\text{Cl}^- \rightarrow \text{Mg}^{2+} + 2\text{Cl}^- + \text{H}_2$

b) $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$

c) $\text{CaCl}_2 + 2\text{Li} \rightarrow 2\text{LiCl} + \text{Ca}$

d) $\text{Ca} + \text{H}_2\text{SO}_4 \rightarrow \text{CaSO}_4 + \text{H}_2$

e) $2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$

f) $\text{MgCl}_2 + \text{F}_2 \rightarrow \text{MgF}_2 + \text{Cl}_2$

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