

Name \_\_\_\_\_

## BONDING



Directions: Use the chart below and your periodic table to answer the questions that follow.

### PART I: COMPARING ELEMENTS

Group One and Group Two both show different compounds made by bonding different elements together. There are two types of bonding in food science, as shown in these two groupings. Let's take a closer look at the elements in each compound to see what makes a compound IONIC or COVALENT.

IONIC BONDS	COVALENT BONDS
NaCl (Sodium Chloride)	H <sub>2</sub> O (Dihydrogen Monoxide)
MgO (Magnesium Oxide)	CH <sub>4</sub> (Methane)
Li <sub>2</sub> O (Lithium Oxide)	CO <sub>2</sub> (Carbon Dioxide)
KF (Potassium Fluoride)	HF (Hydrofluoric Acid)
FeBr <sub>3</sub> (Iron III Bromide)	NH <sub>3</sub> (Ammonia)
CaCl <sub>2</sub> (Calcium Chloride)	NO <sub>2</sub> (Nitrogen Dioxide)
NiI <sub>2</sub> (Nickel Iodide)	C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> (Glucose)
BaS (Barium Sulfide)	CF <sub>4</sub> (Carbon Tetrafluoride)

1. List out the FIRST element in each of the GROUP ONE compounds.
2. What do you notice about the FIRST element that makes up each compound in GROUP ONE?
3. What side of the periodic table do you find these FIRST elements on?
4. List out the SECOND element in each of the GROUP ONE compounds.
5. What do you notice about the SECOND element that makes up each compound in GROUP ONE?
6. What side of the periodic table do you find these SECOND elements on?

**BIG IDEA #1: When a \_\_\_\_\_ and a \_\_\_\_\_ come together, they make an IONIC BOND.**

7. What do you notice about ALL of the elements in GROUP TWO?
8. What side of the periodic table do you find ALL of these elements on?

# Ionic Bonding Pogil Answers

**M. S. Sethi, M. Satake**



## **Ionic Bonding Pogil Answers:**

Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more      **Ionic Bonds | Characteristics of Ionic Bonds and Properties of Ionic Compounds | Grade 6-8 Physical Science** Dot EDU, 2024-04-15 Dive into the electrifying world of Ionic Bonds Characteristics and Properties for grades 6 8 This insightful book unravels the mysteries of ionic bonds and compounds explaining their formation characteristics and role in the matter s structure Discover the intricacies of atoms the exchange of electrons and the fascinating properties of ionic compounds through engaging examples and explanations Perfect for educators and students this resource illuminates a foundational concept in physical science Add it to your science curriculum to spark curiosity and understanding in young learners      Chemical Bonding Kok Koon Leong, 2021-05-03 This book introduces the principles behind chemical bonding to teenagers between the ages of fifteen to seventeen Topics covered include ionic bonding covalent bonding and metallic bonding      **Atoms & Chemical Bonding Science Learning Guide** NewPath Learning, 2014-03-01 The Atoms Atomic Configuration Chemical Bonding Ionic Bonding Ionic Compounds Covalent Bonding Covalent Compounds Naming Compounds and Metallic Bonding Aligned to Next Generation Science Standards NGSS and other state standards      **Teaching Chemical Bonding** Margaret Irene Lindsay, 1995 This document presents an instructional strategy for teaching chemical bonding using parables and music Games student interactions and worksheets are included in the lesson plans Topics include metallic bonding covalent bonding including molecular and network structure and ionic bonding JRH      **Ionic Bonding** , 2006      **Ionic Compounds** Claude H. Yoder, 2007-01-09 A practical introduction to ionic compounds for both mineralogists and chemists this book bridges the two disciplines It explains the fundamental principles of the structure and bonding in minerals and emphasizes the relationship of structure at the atomic level to the symmetry and properties of crystals This is a great reference for those interested in the chemical and crystallographic properties of minerals      **Chemistry** Richard S. Moog, John J. Farrell, 2017-06-26 In the newly updated 7th Edition Chemistry A Guided Inquiry continues to follow the underlying principles developed by years of extensive research on how students learn and draws on testing by those using the POGIL methodology This text follows the principles of inquiry based learning and correspondingly emphasizes underlying chemistry concepts and the reasoning behind them This text provides an approach that follows modern cognitive learning principles by having students learn how to create knowledge based on experimental data and how to test that knowledge      **CHEMICAL BONDING** NARAYAN CHANGDER, 2024-03-31

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*The Chemical Bond in Inorganic Chemistry* Ian David Brown, 2002 there has long been a need for a dedicated monograph on the subject a highly readable book about a theory that though it has long found application in inorganic crystal chemistry deserves to be used more widely Crystallography News The bond valence model is a recently developed model of the chemical bond in inorganic chemistry that complements the bond model widely used in organic chemistry It is simple quantitative intuitive and predictive no more than a pocket calculator is needed to calculate it This book focuses on the theory that underlies the model and shows how it has been used in physics materials science chemistry mineralogy soil science and molecular biology **Ionic Bonding Curriculum Unit** Mary Frank Lamar, 2020 *Thermodynamics of Ionic Bonding with Crown Ethers in Aqueous Solution* Stephen W. Adams, 2021 *The Chemical Bond in Inorganic Chemistry* I. David Brown, 2016-09-02 The bond valence model a description of acid base bonding is widely used for analysing and modelling the structures and properties of solids and liquids Unlike other models of inorganic chemical bonding the bond valence model is simple intuitive and predictive and is accessible to anyone with a pocket calculator and a secondary school command of chemistry and physics This new edition of *The Chemical Bond in Inorganic Chemistry The Bond Valence Model* shows how chemical properties arise naturally from the conflict between the constraints of chemistry and those of three dimensional space The book derives the rules of the bond valence model as well as those of the traditional covalent ionic and popular VSEPR models by identifying the chemical bond with the electrostatic flux linking the bonded atoms Most of the new edition is devoted to showing how to apply these ideas to real materials including crystals liquids glasses and surfaces The work includes detailed examples of applications and the final chapter explores the relationship between the flux and quantum

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**Chemical Bonding** Source Wikipedia, 2013-09 Please note that the content of this book primarily consists of articles available from Wikipedia or other free sources online Pages 157 Chapters Covalent bond Hydrogen bond Chemical bond Intermolecular force Electronegativity Ionic bond Pauli exclusion principle Molecular orbital Metallic bond Disulfide bond Atomic orbital Dipolar bond Ligand Peptide bond Van der Waals force Mesomeric effect Electron counting Cooperative binding Covalent radius Aromaticity Hypervalent molecule Halogen bond Quantum mechanical explanation of intermolecular interactions Molecular orbital diagram Radical Hyperconjugation Resonance Polyhedral skeletal electron pair theory Orbital hybridisation Electron affinity Crystal field theory Anodic bonding Ruggero Santilli Salt bridge Non innocent ligand Chemical polarity Lewis structure Valence Valence electron Quintuple bond Coordination geometry Bonding in solids Walsh diagram Hapticity 18 Electron rule Electrovalency Ligand field theory Conjugated system Lennard Jones potential Bent's rule Inert pair effect Atoms in molecules Cation pi interaction Jahn Teller effect Dihydrogen complex Valence bond theory Bent bond Morse potential Hydrophobic effect Stacking Lanthanide contraction Formal charge Octet rule Entropic force London dispersion force Slater's rules Coordination number Metal ligand multiple bond Auophilicity Linear combination of atomic orbitals molecular orbital method Agostic interaction Inductive effect Electron localization function Intercalation Double bond rule Bond valence method Wafer bonding Pyramidal alkene Bound state Catch bond Cooperativity Metal aromaticity Three center four electron bond Sextuple bond Chemical bonding model Antibonding Sigma bond Denticity Pi bond

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