

# 7.2

## IONIC BONDS AND IONIC COMPOUNDS

### Section Review

#### Objectives

1. Explain the formation of ionic compounds.
2. Determine the composition of ionic compounds.

#### Vocabulary

- |                    |                   |
|--------------------|-------------------|
| a. ionic compound  | b. cation         |
| c. anion           | d. ionic compound |
| e. chemical change |                   |

#### Part A: Completion

Write the complete name of each substance, and describe each compound and how it is composed. Write down each compound's chemical formula and name. (Write answers in order.)

1. Sodium and chlorine atoms are needed for sodium chloride. The name of sodium chloride is NaCl. Sodium ions appear in this compound as Na<sup>+</sup>. Chloride ions are negatively charged. Sodium compounds are electrolytes. NaCl has a high melting point and is soluble in water. It is an ionic compound.

2. Magnesium and oxygen atoms form magnesium oxide. The name of magnesium oxide is MgO. Magnesium ions appear in this compound as Mg<sup>2+</sup>. Oxygen ions are negatively charged. Magnesium compounds are electrolytes. MgO has a high melting point and is soluble in water. It is an ionic compound.

#### Part B: True-False

Write true or false next to each statement. (If false, give the correct name.)

- True 1. Sodium and chlorine atoms combine to form sodium chloride. Sodium chloride is a compound.

2. The chemical formula for sodium chloride is NaCl.
3. Sodium ions are Na<sup>+</sup>.
4. Chloride ions are Cl<sup>-</sup>.
5. Sodium chloride is an electrolyte.
6. Sodium chloride has a high melting point.
7. Sodium chloride is soluble in water.
8. Sodium chloride is an ionic compound.
9. Sodium chloride is an electrolyte.
10. Sodium chloride is an ionic compound.

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# Ionic Bonds Chapter 7 Section 7 2 Answers

**M Mark**



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