

Name: Sam Li

Simulation: Ionic and Covalent Bonding

Background

In this investigation you will bond select atoms. Based upon the types of atoms that you choose to combine, you will create either an ionic compound or a covalent compound. You will have the opportunity to analyze the differences between these different types of compounds and to predict the number of atoms needed to create each, as well as learn how to appropriately name them.

1. Describe the difference between an atom and a molecule:

Atoms are single particles, and molecules are made up of multiple atoms

2. Where are metal atoms located on the periodic table? Where are non-metal atoms located on the periodic table?

Metals are on the left side and non-metals are on the right

3. What subatomic particle(s) participate in chemical bonding?

Electron

4. In your own words, define *valence electron*.

electrons on outside ring which allow for bonding, negative charge

5. How can you determine the number of valence electrons in an atom using the periodic table?

Look at the atomic #

6. Draw a Lewis Dot Structure for the following atoms:

a. Strontium (Sr)



b. Carbon (C)



c. Iodine (I)



d. Xenon (Xe)

*Check your answers before moving on to the next portion of the activity.

Ionic Or Covalent Bonding Lab Answers

Victor M. Corman



Ionic Or Covalent Bonding Lab Answers:

Lab Manual Science Class 09 Neena Sinha, R.Rangarajan, Rajesh Kumar, These Lab Manuals provide complete information on all the experiments listed in the latest CBSE syllabus The various objectives materials required procedures inferences etc have been given in a step by step manner Carefully framed MCQs and short answers type questions given at the end of the experiments help the students prepare for viva voce Science Lab Manual Neena Sinha, R Rangarajan, R P Manchanda, R K Gupta, Rajesh Kumar, Lab Manual **Lab Manual Science Class 10** Neena Sinha, R.Rangarajan, Rajesh Kumar, These Lab Manuals provide complete information on all the experiments listed in the latest CBSE syllabus The various objectives materials required procedures inferences etc have been given in a step by step manner Carefully framed MCQs and short answers type questions given at the end of the experiments help the students prepare for viva voce

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Prentice Hall Science Explorer Michael J. Padilla,2002 Chemical Interactions ,2005 **Chemistry** McGraw-Hill Staff,2001-07 **PE Lab Exp(Noncons)Mod Chem 90** Tzimopoulo,1990 *General, Organic, and Biological Chemistry Study Guide and Selected Solutions* Karen C. Timberlake,2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems **Polarity, Solutions, and Separation Science** Kenda Jo Lemont,1998

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illustrations and an emphasis throughout on laboratory safety Again throughout a Consider This feature encourages users to expand the principles of the experiment into interesting applications open ended experiments or unexplored corners Most experiments in the manual can be completed in one lab session but some can be linked or extended for a multi lab project

Lab Manual T/a Human Physiology Bill W. Tillery, 2001-03 *Focus on physical science* Charles H. Heimler, 1989

Chemistry Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02

Glencoe Science McGraw-Hill Staff, 2001-06 *Cooperative Chemistry Lab Manual* Cooper, 2005-02

The laboratory course described in the lab manual emphasizes experimental design data analysis and problem solving Inherent in the design is the emphasis on communication skills both written and oral Students work in groups on open ended projects in which they are given an initial scenario and then asked to investigate a problem There are no formalized instructions and students must plan and carry out their own investigations

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