Stoichiometry Practice 1.1 - Answer Key

Objectives: - Use stouchsometry to calculate the mass of any substance sevelved in a chemical reaction

Directions: Answer the following stoichiometry problems and show your work.

Mans A -> Mole A -> Mole B;

I. When calcium bydrousde and phosphoric acid (H₂PO₄) are combined, calcium phosphate and water are produced. If the reaction starts with 10.0 grams of phosphoric acid, how many moles of water are produced?

Mole A → Mole B → Mass B:

2. How many grams of oxygen are needed in the combustion 1.07 mul of methyl alcohol (CH)OH)?

$$1:07 \text{ mod } \times \left(\frac{1 \text{ mod } O_{A}}{1 \text{ mod } OH_{A}OH}\right) \left(\frac{12.08 \text{ g/O}_{A}}{1 \text{ mod } O_{A}}\right) = 0.1.4 \text{ g}$$

Mans A → Mole A → Mole B → Mass B:

 How many grams of copper are produced when 3.46 grams of aluminum react with excess copper (II) acetate solution? [12.2 g]

 Lithium sulfate and barium chloride react to produce burium sulfate and lithium chloride. How many grams of burium chloride are needed to react with 17.5 g of lithium sulfate?

17.5 g ×
$$\left(\frac{1 \text{ mod } \Delta_1 \otimes O_2}{2 \text{ mod } \Omega_1 \otimes O_2}\right) \left(\frac{1 \text{ mod } ReO_2}{1 \text{ mod } \Omega_1 \otimes O_2}\right) \left(\frac{2 \text{ mod } ReO_2}{1 \text{ mod } ReO_2}\right) = 13.1 g$$

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Y Pai

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