

Molarity, Dilutions, and Preparing Solutions

A. Preparing a Stock Solution of Known Concentration

1. Mass of 100-mL volumetric flask, g 59.45
2. Mass of 100-mL volumetric flask and CuSO_4 , g 64.08
3. Mass of CuSO_4 , g 4.63 g
4. Volume of stock solution, mL 100 mL
5. Molarity of stock solution "S," M 0.0290 M
(Show calculations.)
$$\frac{4.63}{159.6} = 0.0290$$

B. Preparing the Spectrophotometer for use

1. Percent transmission of DI H_2O , %T 100 %
2. Absorbance of DI H_2O , A 0

C. Measuring the Absorbance of the Stock Solution

1. Percent transmission of stock solution, %T 43.9
2. Absorbance of stock solution, A 0.357

E. Construction of a Calibration Graph

	"A"	"B"	"C"	"D"
Volume of stock solution, mL	5.00	10.00	15.00	20.00
Volume of prepared solution, mL	50.00	50.00	50.00	50.00
Molarity, M	<u>0.0058</u>			
Percent transmission, %T	<u>105</u>	<u>82.3</u>	<u>79.5</u>	<u>52.6</u>
Absorbance, A	<u>.021</u>	<u>.085</u>	<u>.100</u>	<u>.280</u>

Molarity Concentration Lab Report

DP Hallahan

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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