REDOX REACTIONS

Name ____

For the equations below, identify the substance oxidized, the substance reduced, the oxidizing agent, the reducing agent, and write the oxidation and reduction half reactions.

1. 2H, + 0, - 2H,0 OX: 2H, -> 4H + + 4e RED: 03° + 4e -> 20-2.

loducing agent

2. Fo + $2n^{2}$ + Fe^{2} + Zn $0x: Fe \longrightarrow Fe^{2}$ + Ze^{-} $red: Zn^{+2} + Ze^{-} \longrightarrow Zn^{0}$

reducing agent

3. 2AI + 3Fe⁻¹² -> 2AI⁻³ + 3Fe $0x : \partial AI^{\circ} -> \partial AI^{-13} + 6e^{-}$ reducing agent

Red: $3Fe^{+2} + 6e^{-} -> 3Fe$ oxidizing agent

4. Cu + 2AgNO, - CU(NO), + 2Ag

Ox: Cu° -> Cu+2 + 2e reducing agent

Ted: 2Ag+ + 2e -> 2Ag oxidizing agent

NO2 = Spectator ion

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Bill Linderman

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