

2. A chemist measures the volume of a gas at various pressures. The volumes she measured are in the following table next to each pressure that was tested.

Pressure (atm)	Volume (mL)
1.00	4.02
1.30	3.10
1.60	2.51
1.90	2.11
2.20	1.84

- a. Pressure and volume are NOT directly proportional as in the previous problem. Instead they are *inversely proportional*. To make a linear graph, first create a new column of data containing the inverse of the pressure values ( $1/P$ ).

$1/P$ ( )

- b. Graph  $1/P$  vs. volume, format the graph, and include the best-fit line with the equation.
- c. Calculate the volume of the balloon when the pressure is adjusted to 1.50 atm.

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**Siegel**



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