

Reaction Rates and Chemical Equilibrium

A. Factors That Affect the Rate of a Reaction

A.1 Effect of Temperature

Test Tube Temperature (3)

Observations (4)

1. 10 °C

Slow fizzing, clear solution after 2 min 38 sec

2. 58 °C

vigorous fizzing/bubbles, clear solution in

Which test tube cleared first (4)?

48 sec

Q1 How did an increase in temperature affect the rate of a reaction?

A.2 Effect of Changing Reactant Concentration

Test Tube	Concentration of HCl	Initial time (2)	Final time (3)	Total time (3) - (2)
1.	1.0 M HCl	0	17 min 5 sec	
2.	2.0 M HCl	0	4 min 38 sec	
3.	3.0 M HCl	0	1 min 20 sec	

Rank the test tubes in order of slowest to fastest. _____ < _____ < _____

Q2 How did an increase in the concentration of HCl cause a difference in the rates of reaction of HCl with Mg?

Lab Report On Reaction Rates And Equilibrium

Sabine Zange



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Reaction Rates for High-temperature Air with Carbon and Sodium Impurities Mina L. Carnicom, 1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have been compiled and are presented in tabular form Included are reactions involving neutral species charged species free electrons some species in excited electronic or vibrational states and radiative processes

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Catalog University of Maine at Orono, 1929 Evaluated Kinetic Data for High Temperature Reactions D. L. Baulch, 1972

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