#### 1. (4 points)

The specific heat (s) of iron (Fe) is 0.40 J/(g K). Calculate the heat effect q when 20 g of iron is heated from 100 °C to 200 °C. Include the sign!

 $q = m.s.\Delta T$  for  $\Delta T \Delta K = \Delta^{\circ}C$  q = 20x0.40x(200-100) = +800 J

You must include the units and sign in the answer!

#### 2. (4 points)

Reaction:  $Mg(s) + \frac{1}{2}O_3(g) \rightarrow MgO(s)$   $\Delta H = -600 \text{ kJ/mol}$ 

The Molar Mass of Mg = 24.0 g/mol.

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

 $2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol})x(-600)(kJ/\text{mol}) = -60 \text{ kJ}$ 

Again, your answer must include both the sign and the units.

 (2 points) Which one of the following four reactions is a formation reaction? <u>Circle</u> the correct answer.

a) 
$$H_2O(I) + \frac{1}{2}O_2 \longrightarrow H_2O_2(I)$$

b) 
$$SO_2(g) + \frac{1}{2}O_2 \rightarrow SO_2(g)$$

c) 
$$Cl(g) + H(g) \rightarrow HCl(g)$$

d) 
$$S(s) + O_2(g) \rightarrow SO_2(g)$$

Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur, so e.g.  $Cl_2$  not  $Cl_4$  and  $H_2$  not  $H_3$ !

#### CHEM 101 Section L52 Quiz #6 b

#### 1. (4 points)

The specific heat (s) of copper (Cu) is 0.38 J/(g.K). Calculate the heat effect q when 10 g of copper is heated from 100 °C to 300 °C. Include the sign!

 $q = m.s.\Delta T$  for  $\Delta T.\Delta K = \Delta^*C$  q = 10x0.38x(300-100) = +760 J

You must include the units and sign in the answer!

#### 2. (4 points)

Reaction:  $Ca(s) + \frac{1}{2}O_2(g) \rightarrow CaO(s)$ 

$$\Delta H = -650 \text{ kJ/mol}$$

The Molar Mass of Ca = 40.0 g/mol.

Calculate the enthalpy change  $\Delta H$  when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

 $8.0 \text{ g Ca} = 8.0/40/ = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol})x(-650)(kJ/\text{mol}) = -130 \text{ kJ}$ 

Again, your answer must include both the sign and the units.

 (2 points) Which one of the following reactions is a formation reaction? <u>Circle</u> the correct answer.

a) 
$$Br(g) + H(g) \rightarrow HBr(g)$$

b) 
$$H_2O(I) + \frac{1}{2}O_2 \rightarrow H_2O_2(I)$$

c) 
$$2P(s) + 2\%O_2(g) \rightarrow P_2O_3(s)$$

d) 
$$SO_2(g) + \frac{1}{2}O_2 \rightarrow SO_3(g)$$

# **Merrill Chemistry Chapter 8 Review Answers**

Robert C. Smoot, Smoot, Richard G. Smith, Jack Price

## **Merrill Chemistry Chapter 8 Review Answers:**

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Merrill Earth Science Ralph M. Feather, Susan Leach Snyder, Dale T. Hesser, 1995 Physical Chemistry of Metallurgical Processes, Second Edition Mohammad Shamsuddin, 2021-06-22 This updated second edition retains its classroom tested treatment of physical chemistry of metallurgical topics such as roasting of sulfide minerals matte smelting converting structure properties and theories of slag reduction of oxides and reduction smelting interfacial phenomena steelmaking secondary steelmaking role of halides in extraction of metals refining hydrometallurgy and electrometallurgy and adds new data in worked out examples as well as up to date references to the literature The book further explains the physical chemistry of various metallurgical topics steps involved in extraction of metals such as roasting matte smelting converting reduction smelting steelmaking reactions deoxidation stainless steelmaking vacuum degassing refining leaching chemical precipitation ion exchange solvent extraction cementation gaseous reduction and electrowinning Each topic is illustrated with appropriate examples of applications of the technique in extraction of some common reactive rare or refractory metal together with worked out problems explaining the principle of the operation The problems require imagination and critical analyses and also encourage readers for creative application of thermodynamic data in metal extraction Updates and condenses text throughout the book by sequential arrangement of paragraphs in different chapters Maximizes readers understanding of the physicochemical principles involved in extraction production of common and rare reactive metals by pyro as well as hydrometallurgical routes Reinforces concepts presented with worked examples in each chapter explaining the process steps Explains the physical chemistry of various metallurgical steps such as roasting matte smelting converting and reduction smelting steelmaking aqueous processing etc in extraction of metals Collects and uniformly presents scattered information on physicochemical principles of metal production from various books and journals **Technical Book Review**, 1963 El-Hi Textbooks and Natural Inorganic Hydrochemistry in Relation to Serials in Print, 1995 R R Bowker Publishing, 1995-02 **Groundwater** J. W. Lloyd, J. A. Heathcote, 1985 Provides comprehensive but concise coverage of groundwater chemistry theory hydrochemical processes and methods of interpretation It has particular relevance in the study of pollution problems **Polymorphism in Pharmaceutical Solids** Harry G. Brittain, 2018-11-12 Using clear and practical examples

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