

Part A

Answer all questions in this part.

Directions (1–30): For each statement or question, record on your separate answer sheet the *number* of the word or expression that, of those given, best completes the statement or answers the question. Some questions may require the use of the 2011 Edition Reference Tables for Physical Setting/Chemistry.

1 Which statement describes the charge of an electron and the charge of a proton?

- (1) An electron and a proton both have a charge of +1.
- (2) An electron and a proton both have a charge of -1.
- (3) An electron has a charge of +1, and a proton has a charge of -1.
- (4) An electron has a charge of -1, and a proton has a charge of +1.

2 Which subatomic particles are found in the nucleus of an atom of beryllium?

- (1) electrons and protons
- (2) electrons and positrons
- (3) neutrons and protons
- (4) neutrons and electrons

3 The elements in Period 4 on the Periodic Table are arranged in order of increasing

- (1) atomic radius
- (2) atomic number
- (3) number of valence electrons
- (4) number of occupied shells of electrons

4 Which phrase describes two forms of solid carbon, diamond and graphite, at STP?

- (1) the same crystal structure and the same properties
- (2) the same crystal structure and different properties
- (3) different crystal structures and the same properties
- (4) different crystal structures and different properties

5 Which element has six valence electrons in each of its atoms in the ground state?

- (1) Se
- (2) As
- (3) Kr
- (4) Ga

6 What is the chemical name for $\text{H}_2\text{SO}_3(\text{aq})$?

- (1) sulfuric acid
- (2) sulfurous acid
- (3) hydrosulfuric acid
- (4) hydrosulfurous acid

7 Which substance is most soluble in water?

- (1) $(\text{NH}_4)_3\text{PO}_4$
- (2) $\text{Cu}(\text{OH})_2$
- (3) Ag_2SO_4
- (4) CaCO_3

8 Which type of bonding is present in a sample of an element that is malleable?

- (1) ionic
- (2) metallic
- (3) nonpolar covalent
- (4) polar covalent

9 Which atom has the greatest attraction for the electrons in a chemical bond?

- (1) hydrogen
- (2) oxygen
- (3) silicon
- (4) sulfur

10 Which type of reaction involves the transfer of electrons?

- (1) alpha decay
- (2) double replacement
- (3) neutralization
- (4) oxidation-reduction

11 A 10.0-gram sample of nitrogen is at STP. Which property will increase when the sample is cooled to 72 K at standard pressure?

- (1) mass
- (2) volume
- (3) density
- (4) temperature

12 Which element is a gas at STP?

- (1) sulfur
- (2) xenon
- (3) potassium
- (4) phosphorus

June 2003 Physics Regents Answer Key

**Samlal Mannie, Caribbean
Examinations Council**



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Table of Contents June 2003 Physics Regents Answer Key

1. Understanding the eBook June 2003 Physics Regents Answer Key
 - The Rise of Digital Reading June 2003 Physics Regents Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying June 2003 Physics Regents Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an June 2003 Physics Regents Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from June 2003 Physics Regents Answer Key
 - Personalized Recommendations
 - June 2003 Physics Regents Answer Key User Reviews and Ratings
 - June 2003 Physics Regents Answer Key and Bestseller Lists
5. Accessing June 2003 Physics Regents Answer Key Free and Paid eBooks

- June 2003 Physics Regents Answer Key Public Domain eBooks
 - June 2003 Physics Regents Answer Key eBook Subscription Services
 - June 2003 Physics Regents Answer Key Budget-Friendly Options
6. Navigating June 2003 Physics Regents Answer Key eBook Formats
- ePub, PDF, MOBI, and More
 - June 2003 Physics Regents Answer Key Compatibility with Devices
 - June 2003 Physics Regents Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of June 2003 Physics Regents Answer Key
 - Highlighting and Note-Taking June 2003 Physics Regents Answer Key
 - Interactive Elements June 2003 Physics Regents Answer Key
8. Staying Engaged with June 2003 Physics Regents Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers June 2003 Physics Regents Answer Key
9. Balancing eBooks and Physical Books June 2003 Physics Regents Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection June 2003 Physics Regents Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine June 2003 Physics Regents Answer Key
- Setting Reading Goals June 2003 Physics Regents Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of June 2003 Physics Regents Answer Key
- Fact-Checking eBook Content of June 2003 Physics Regents Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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