Stoichiometry

Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write true or false.

- The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
- Steichiometry is based on the law of conservation of mass.
- In any chemical reaction, the mass of the products is less than the mass of the reactants.
- 4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
 - The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.

methanol + oxygen → carbon dioxide + water

 $2\mathrm{CH_3OH(l)} + 3\mathrm{O_2}\left(g\right) \rightarrow 2\mathrm{CO_2}(g) + 4\mathrm{H_2O}(g)$

Substance	Molar Mass (g/mol)	Mumber of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05		2	
7. Oxygen gas	32.00			
8. Carbon dioxide	44.01		(i)	
9. Water	18.02			

- 10. What are the reactants?
- 11. What are the products?
- 12. What is the total mass of the reactants?
- 13. What is the total mass of the products?
- 14. How do the total masses of the reactants and products compare?

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Table of Contents Mcgraw Hill Night Study Guide Teacher Guide

- 1. Understanding the eBook Mcgraw Hill Night Study Guide Teacher Guide
 - The Rise of Digital Reading Mcgraw Hill Night Study Guide Teacher Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Mcgraw Hill Night Study Guide Teacher Guide
 - \circ Exploring Different Genres
 - $\circ\,$ Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Night Study Guide Teacher Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Mcgraw Hill Night Study Guide Teacher Guide
 - Personalized Recommendations
 - Mcgraw Hill Night Study Guide Teacher Guide User Reviews and Ratings
 - Mcgraw Hill Night Study Guide Teacher Guide and Bestseller Lists
- 5. Accessing Mcgraw Hill Night Study Guide Teacher Guide Free and Paid eBooks
 - Mcgraw Hill Night Study Guide Teacher Guide Public Domain eBooks
 - Mcgraw Hill Night Study Guide Teacher Guide eBook Subscription Services
 - Mcgraw Hill Night Study Guide Teacher Guide Budget-Friendly Options
- 6. Navigating Mcgraw Hill Night Study Guide Teacher Guide eBook Formats
 - o ePub, PDF, MOBI, and More
 - Mcgraw Hill Night Study Guide Teacher Guide Compatibility with Devices
 - Mcgraw Hill Night Study Guide Teacher Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Mcgraw Hill Night Study Guide Teacher Guide
 - Highlighting and Note-Taking Mcgraw Hill Night Study Guide Teacher Guide
 - o Interactive Elements Mcgraw Hill Night Study Guide Teacher Guide
- 8. Staying Engaged with Mcgraw Hill Night Study Guide Teacher Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Mcgraw Hill Night Study Guide Teacher Guide
- 9. Balancing eBooks and Physical Books Mcgraw Hill Night Study Guide Teacher Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Night Study Guide Teacher Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Hill Night Study Guide Teacher Guide
 - Setting Reading Goals Mcgraw Hill Night Study Guide Teacher Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Hill Night Study Guide Teacher Guide
 - Fact-Checking eBook Content of Mcgraw Hill Night Study Guide Teacher Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - o Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - $\circ \ \ Integration \ of \ Multimedia \ Elements$
 - Interactive and Gamified eBooks

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