

CHAPTER **11** STUDY GUIDE

# Stoichiometry

## Section 11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the balanced equation.

For each statement below, write *true* or *false*.

- \_\_\_\_\_ 1. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction is called stoichiometry.
- \_\_\_\_\_ 2. Stoichiometry is based on the law of conservation of mass.
- \_\_\_\_\_ 3. In any chemical reaction, the mass of the products is less than the mass of the reactants.
- \_\_\_\_\_ 4. The coefficients in a chemical equation represent not only the number of individual particles but also the number of moles of particles.
- \_\_\_\_\_ 5. The mass of each reactant and product is related to its coefficient in the balanced chemical equation for the reaction by its molar mass.

Complete the table below, using information represented in the chemical equation for the combustion of methanol, an alcohol.



Substance	Molar Mass (g/mol)	Number of Molecules	Number of Moles (mol)	Mass (g)
6. Methanol	32.05			
7. Oxygen gas	32.00			
8. Carbon dioxide	44.01			
9. Water	18.02			

10. What are the reactants? \_\_\_\_\_
11. What are the products? \_\_\_\_\_
12. What is the total mass of the reactants? \_\_\_\_\_
13. What is the total mass of the products? \_\_\_\_\_
14. How do the total masses of the reactants and products compare? \_\_\_\_\_

# Mcgraw Study Guide

**G Thomas**



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## **Table of Contents Mcgraw Study Guide**

1. Understanding the eBook Mcgraw Study Guide
  - The Rise of Digital Reading Mcgraw Study Guide
  - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Mcgraw Study Guide
  - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Study Guide
  - Personalized Recommendations
  - Mcgraw Study Guide User Reviews and Ratings
  - Mcgraw Study Guide and Bestseller Lists
5. Accessing Mcgraw Study Guide Free and Paid eBooks
  - Mcgraw Study Guide Public Domain eBooks
  - Mcgraw Study Guide eBook Subscription Services
  - Mcgraw Study Guide Budget-Friendly Options
6. Navigating Mcgraw Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- Mcgraw Study Guide Compatibility with Devices
- Mcgraw Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Mcgraw Study Guide
  - Highlighting and Note-Taking Mcgraw Study Guide
  - Interactive Elements Mcgraw Study Guide
- 8. Staying Engaged with Mcgraw Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Mcgraw Study Guide
- 9. Balancing eBooks and Physical Books Mcgraw Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Mcgraw Study Guide
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Mcgraw Study Guide
  - Setting Reading Goals Mcgraw Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Mcgraw Study Guide
  - Fact-Checking eBook Content of Mcgraw Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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