

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER 12

## STUDY GUIDE FOR CONTENT MASTERY

## Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of  $\text{O}_2$ ,  $\text{N}_2$ , and  $\text{NO}_2$  are 32.00 g/mol, 28.02 g/mol, and 46.01 g/mol, respectively.

Amount of $\text{O}_2$	Amount of $\text{N}_2$	Amount of $\text{NO}_2$	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	$\text{N}_2$	2 molecules $\text{O}_2$
2 molecules	4 molecules	4 molecules	$\text{O}_2$	2 molecules $\text{N}_2$
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	4.00 mol	4.00 mol	$\text{N}_2$	2.00 mol $\text{O}_2$
2.00 mol	4.00 mol	4.00 mol	$\text{O}_2$	2.00 mol $\text{N}_2$
1.00 mol	4.00 mol	4.00 mol	$\text{O}_2$	2.00 mol $\text{N}_2$
0.500 mol	0.200 mol	0.200 mol	$\text{N}_2$	0.300 mol $\text{O}_2$
32.00 g	56.04 g	92.02 g	none	none
64.00 g	56.04 g	92.02 g	$\text{O}_2$	32.00 g $\text{N}_2$
64.00 g	112.08 g	184.04 g	$\text{O}_2$	48.01 g $\text{N}_2$

balanced equation mole ratio = 2 mol  $\text{NO}_2$ /1 mol  $\text{O}_2$

16.00 g  $\text{O}_2$   $\times$  1 mol  $\text{O}_2$ /32.00 g  $\text{O}_2$  = 0.500 mol  $\text{O}_2$

28.02 g  $\text{N}_2$   $\times$  1 mol  $\text{N}_2$ /28.02 g  $\text{N}_2$  = 1.000 mol  $\text{N}_2$

actual mole ratio = 0.500 mol  $\text{N}_2$ /0.500 mol  $\text{O}_2$  = 1.000 mol  $\text{N}_2$ /1.000 mol  $\text{O}_2$

Because the actual mole ratio of  $\text{N}_2$ / $\text{O}_2$  is larger than the balanced equation mole ratio of  $\text{N}_2$ / $\text{O}_2$ , there is an excess of  $\text{N}_2$ ;  $\text{O}_2$  is the limiting reactant.

Mass of  $\text{NO}_2$  used = 0.500 mol  $\text{O}_2$   $\times$  2 mol  $\text{NO}_2$ /1 mol  $\text{O}_2$  = 1.000 mol  $\text{NO}_2$

0.500 mol  $\text{NO}_2$   $\times$  46.01 g  $\text{NO}_2$ /1 mol  $\text{NO}_2$  = 23.01 g  $\text{NO}_2$

Mass of  $\text{NO}_2$  produced = 0.500 mol  $\text{NO}_2$   $\times$  46.01 g  $\text{NO}_2$ /1 mol  $\text{NO}_2$  = 23.01 g  $\text{NO}_2$

Excess  $\text{N}_2$  = 28.02 g  $\text{N}_2$  - 23.01 g  $\text{N}_2$  = 5.01 g  $\text{N}_2$

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## CHAPTER 13

## STUDY GUIDE FOR CONTENT MASTERY

## Section 13.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



**Example Problem:** The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and  $\text{Ga}_2\text{O}_3$  are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown 7.42 g  $\text{Ga}_2\text{O}_3$
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: 7.00 g Ga  $\times$   $\frac{69.72 \text{ g Ga}}{1 \text{ mol Ga}}$  = 1 mol Ga/69.72 g Ga
- number of moles of product: 0.100 mol Ga  $\times$   $\frac{2 \text{ mol Ga}_2\text{O}_3}{1 \text{ mol Ga}}$  = 0.200 mol  $\text{Ga}_2\text{O}_3$
- theoretical yield: 0.200 mol  $\text{Ga}_2\text{O}_3$   $\times$   $\frac{157.44 \text{ g Ga}_2\text{O}_3}{1 \text{ mol Ga}_2\text{O}_3}$  = 31.49 g  $\text{Ga}_2\text{O}_3$  correct
- percent yield:  $\frac{7.42 \text{ g Ga}_2\text{O}_3}{31.49 \text{ g Ga}_2\text{O}_3} \times 100$  = 23.6% correct

# Mcgraw Hill Chemistry Answer Key Ch8

**Richard H. Langley, Richard Langley**



## **McGraw Hill Chemistry Answer Key Ch8:**

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