

PHYSICAL CHEMISTRY

(Complete Syllabus)

SECTION - A

Straight Objective Type

This section contains 220 multiple choice questions. Each question has 4 choices (1), (2), (3) and (4), out of which **ONLY ONE** is correct.

Choose the correct answer :

- The mass of oxygen required to produce enough CO which completely reduces x gm of Fe_2O_3 . Suppose molecular weight of $\text{O}_2 = y$, molecular weight of $\text{Fe}_2\text{O}_3 = z$
(1) $\frac{2xz}{3y}$ (2) $\frac{3xy}{2z}$
(3) $\frac{2y}{3xz}$ (4) $\frac{3z}{2xy}$
- The number of lone pair of electrons present in 2.56 gm of orthorhombic sulphur will be
(1) 0.02 N_0 (2) 0.16 N_0
(3) 0.32 N_0 (4) 0.08 N_0
- A 10.0 g of sample of mixture of calcium chloride and sodium chloride is treated with Na_2CO_3 to precipitate the calcium as calcium carbonate. This CaCO_3 is heated to convert all calcium to CaO and final mass of CaO is 1.62 gms. The % by mass of CaCl_2 in the original mixture is
(1) 15.2%
(2) 32.1%
(3) 21.8%
(4) 11.07%
- 1000 ml of oxygen at STP were passed through an ozoniser and the resulting volume was 888 ml at STP. This quantity of ozonised oxygen is passed through excess of KI solution. The weight of I_2 liberated is
(1) 25.4 g (2) 12.7 g
(3) 2.54 g (4) 1.27 g
- x_1 ml solution having normality 1 is diluted to get a solution of normality y_2 . What would be volume required for dilution?
(1) $x_1 \left(1 - \frac{y_1}{y_2}\right)$ (2) $x_1 \left(\frac{y_2}{y_1} - 1\right)$
(3) $x_1 \left(\frac{y_1}{y_2} - 1\right)$ (4) $x_1 \left(1 + \frac{y_1}{y_2}\right)$
- 1.625 g of an interhalogen compound ICl_x was subjected to electric discharge which split it into constituents and brought to STP. It occupied 112 ml. The interhalogen compound is
(1) ICl (2) ICl_3
(3) ICl_5 (4) ICl_7
- If 12.25 g of a sample of KClO_3 of 75% purity and rest vapourisable impurities is strongly heated to give KCl , KClO_4 and 0.32 gm O_2 . The mass of residue obtained is
(1) 1.769 g (2) 7.096 g
(3) 5.327 g (4) 8.85 g
- A 2 g sample of xenon reacts with fluorine. The mass of compound produced is 3.158 g. The empirical formula of the compound is
(1) XeF_2 (2) XeF_4
(3) XeF_6 (4) XeF_8
- 1 g of carbonate of metal was dissolved in 25 ml of 1 N HCl. The resulting liquid required 5 ml of 1 N NaOH for neutralisation. The equivalent weight

Multiple Choice Physical Chemistry Exam Questions

IM Harris



Multiple Choice Physical Chemistry Exam Questions:

Reviewing **Multiple Choice Physical Chemistry Exam Questions**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Multiple Choice Physical Chemistry Exam Questions**," an enthralling opus penned by a highly acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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