

EQUILIBRIUM TEST

MULTIPLE CHOICE QUESTIONS

8) Consider the following reaction:



Which one of the following statements is correct?

- A. Since very large, the reaction is irreversible.
- B. Changing the pressure will affect the value of K_c .
- C. It is an very large, the reaction is effectively irreversible.
- D. Changing the temperature will not affect the value of K_c .

9) The following reaction, in which nitrogen trioxide, N_2O_3 , is converted to nitrogen dioxide, NO_2 , occurs at equilibrium:



In a given closed system, the equilibrium constant for this reaction is 0.50. If K_c and the initial concentration of N_2O_3 at equilibrium is 0.2000 M, at this temperature, the initial concentrations of NO_2 at equilibrium is

- A. 0.2000 M
- B. 1.000 M
- C. 1.40 M
- D. 0.00 M

10) Chloroacetic acid, ClCH_2COOH , is known with water, H_2O , as being weakly acidic. $\text{H}_2\text{ClCH}_2\text{COOH}$, $\text{H}_2\text{ClCH}_2\text{COO}^-$ equilibrium exist with H_2O as follows: $\text{H}_2\text{ClCH}_2\text{COOH}$ and hydrochloric acid, H_2Cl^+ according to the equilibrium below:



It is known that constant 0.1 M of chloroacetic acid is placed in a closed and left overnight and that three hydrochloric acids reaches equilibrium. In this scenario, after change is made in the system, which one of the following statements is most likely to result in an increase in the concentrations of H_3O^+ ions?

- A. Change the solution.
- B. Bubble in more CO_2 gas.
- C. Add a few drops of phenolphthalein indicator.
- D. Add a few drops of 0.1 M sodium hydroxide, NaOH , solution.

Multiple Choice Concept Questions On Equilibrium

Patrick Vollmar



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