

JIPMER-02-06-2019-Morning
Chemistry

1. $^{19}\text{F}^{-1}$, $^{16}\text{O}^{-2}$, ^{20}Ne choose the correct statement
 - (1) Both O^{-2} and F^{-} are isoelectronic.
 - (2) All given have equal no of e^{-}
 - (3) F^{-} and Ne have equal number of e^{-}
 - (4) All

2. Hund's Rule state that:
 - (1) Number of two e^{-} can be in two separate orbitals
 - (2) Number of two e^{-} can be present with similar spin in a orbital.
 - (3) No one e^{-} can exist in 'd' orbital
 - (4) None of these

3. If two atoms have equal number of electron it is called:
 - (1) Iso-electronic
 - (2) isotone
 - (3) isobar
 - (4) None of these

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