

# DE-100A: STUDY GUIDE

**Introduction and Overview** - what topics will you be learning  
 - understanding basic thermodynamics + 1st, 2nd, 3rd laws of thermodynamics  
 - **Energy**  
 - **Enthalpy**  
 - **Entropy**  
 - **Free Energy**  
 - **Equilibrium**  
 - **Chemical Equilibrium**  
 - **Le Chatelier's Principle**



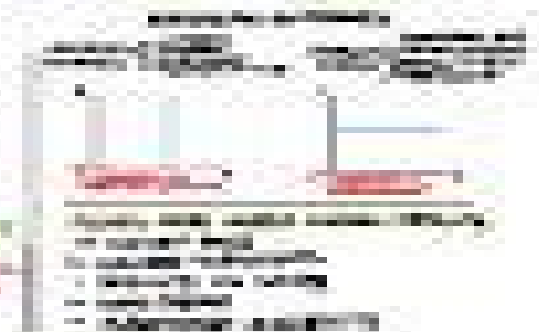
**1st Law of Thermodynamics**  
 - Energy cannot be created or destroyed, only transferred or transformed.  
 -  $\Delta U = q + w$   
 -  $q$  = heat added to the system  
 -  $w$  = work done on the system  
 -  $\Delta U$  = change in internal energy



**2nd Law of Thermodynamics**  
 - The total entropy of an isolated system can never decrease over time.  
 -  $\Delta S_{total} \geq 0$   
 -  $\Delta S_{total} = \Delta S_{system} + \Delta S_{surroundings}$   
 -  $\Delta S_{system} = \frac{q_{rev}}{T}$   
 -  $\Delta S_{surroundings} = -\frac{q_{rev}}{T}$



**3rd Law of Thermodynamics**  
 - The entropy of a perfect crystal is zero at absolute zero temperature.  
 -  $S \rightarrow 0$  as  $T \rightarrow 0$   
 -  $\Delta S = nR \ln \frac{V_2}{V_1}$   
 -  $\Delta S = nR \ln \frac{P_1}{P_2}$



**Chemical Equilibrium**  
 -  $K_{eq} = \frac{[products]}{[reactants]}$   
 -  $\Delta G = -RT \ln K_{eq}$

**Le Chatelier's Principle**  
 - If a system at equilibrium is subjected to a change in concentration, pressure, or temperature, the system will shift to counteract the change.

**Thermodynamic Data**  
 - Standard enthalpy of formation ( $\Delta H_f^\circ$ )  
 - Standard entropy ( $S^\circ$ )  
 - Standard Gibbs free energy of formation ( $\Delta G_f^\circ$ )



# Microeconomics Study Guide Final

**Michael Hoelscher**



## **Microeconomics Study Guide Final:**

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