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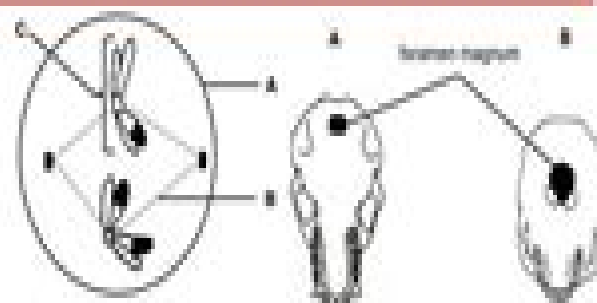
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Life Science Memorandum November 2014

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations
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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction
therefore the net ionic equation is $2\text{cl aq} + 2\text{ag aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you
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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 $2\text{Ag}^+_{\text{aq}} + 2\text{F}^-_{\text{aq}} + 2\text{NH}_4^+_{\text{aq}} + \text{Cr}_2\text{O}_7^{2-}_{\text{aq}} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7\text{s} + 2\text{NH}_4^+_{\text{aq}} + 2\text{F}^-_{\text{aq}}$ because two $\text{NH}_4^+_{\text{aq}}$

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