

11. For  $\Delta_f H^\circ$  reaction, the increasing order of the reactivity of the following alkyl halides is

- (A)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{Br}$   
(B)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{I}$   
(C)  $\text{CH}_3\text{CH}_2\text{Br}$   
(D)  $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{Br}$

Choose the correct answer from the options given below

- (1)  $(A) < (B) < (C) < (D)$   
(2)  $(B) < (A) < (D) < (C)$   
(3)  $(A) < (C) < (B) < (D)$   
(4)  $(C) < (B) < (D) < (A)$

Read the following passage and answer the next five questions based on it.

Battery or cell converts chemical energy of the redox reaction to electrical energy. In fuel cell (a galvanic cell), the chemical energy of combustion of fuels like  $\text{H}_2$ , ethanol, etc. are directly converted to electrical energy. In a fuel cell,  $\text{H}_2$  and  $\text{O}_2$  react to produce electricity, where  $\text{H}_2$  gas is oxidised at anode and oxygen is reduced at cathode and the reactions involved are



87.2 L of  $\text{H}_2$  at STP reacts in 15 minutes.

12. The number of moles of hydrogen oxidised is

- (1) 8.12 moles (2) 13.3 moles (3) 3.8 moles (4) 1.12 moles

13. The number of moles of electrons produced in the oxidation of 87.2 L of  $\text{H}_2$  at STP is

- (1) 2 moles (2) 4 moles (3) 1 mole (4) 8 moles

14. The quantity of electricity produced in the oxidation of 87.2 L of  $\text{H}_2$  at STP is

- (1) 96500 C (2) 193000 C (3) 183000 C (4) 40250 C

15. If the entire current produced is used for the electroplating of silver (at wt.  $108 \text{ g mol}^{-1}$ ) from Silver (I) solution, the amount of silver deposited will be

- (1) 514 g (2) 444 g (3) 108 g (4) 216 g

16. The source of electrical energy in the Apollo moon flight was

- (1) Lead storage battery (2) A galvanic cell  
(3) Ni-Cadmium (4)  $\text{H}_2\text{O}_2$  fuel cell

Read the following passage and answer the next five questions based on it.

Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn
Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd
La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg

In any transition series, as we move from left to right the  $d$ -orbitals are progressively filled and their properties vary accordingly.

Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr			

The above are the two series of  $d$ -block elements in which the chemical properties won't change much. The  $f$ -block elements are radioactive in nature and mostly are artificially synthesised in laboratories and thus much is not known about their chemical properties.

17. Identify the incorrect statement.

- (1) Second ionisation enthalpy of  $\text{Ag}$  is greater than second ionisation enthalpy of  $\text{Pd}$ .  
(2)  $\text{Zr}$  and  $\text{Hf}$  shows almost identical metallic properties.  
(3) Melting point of  $\text{Mn}$  is lower than that of  $\text{Cr}$ .  
(4) Interstitial compounds are non-metallic in nature and neither were nor conductive in nature.

18. Which of the following is the correct order of second ionisation enthalpy?

- (1)  $\text{V} < \text{Cr} < \text{Mn}$  (2)  $\text{V} < \text{Cr} < \text{Mn}$  (3)  $\text{V} < \text{Cr} < \text{Mn}$  (4)  $\text{V} < \text{Cr} < \text{Mn}$

19. Which of the following pair of compounds exhibits same colour in aqueous solution?

- (1)  $\text{FeCl}_3$ ,  $\text{CuCl}_2$  (2)  $\text{VOCl}_3$ ,  $\text{CuCl}_2$   
(3)  $\text{VOCl}_3$ ,  $\text{FeCl}_3$  (4)  $\text{VOCl}_3$ ,  $\text{MnCl}_2$

20. Which metal has the highest oxidation state in the first row transition series?

- (1) Cr (2) Fe (3) Mn (4) V

21. Why do the actinoids exhibit higher number of oxidation states than lanthanoids?

- (1)  $4f$  orbitals are more diffused than the  $5f$  orbitals.  
(2) Energy difference between  $5f$  and  $6d$  is high with respect to the energy difference between  $4f$  and  $5d$ .  
(3) Energy difference between  $5f$  and  $6d$  is lower with respect to the energy difference between  $4f$  and  $5d$ .  
(4) Actinoids are more reactive in nature than the lanthanoids.

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