

- (vi) Some candidates incorrectly answered as efflorescent or hygroscopic.
- (f) (i) Common error that was repeated among all the sub parts was that even if the reagent was correctly chosen the result with both the substances under consideration was not specified.  
The reagent, conc.  $\text{H}_2\text{SO}_4$ , was incorrectly chosen by many candidates.  
The solution state of  $\text{NaCl}$  and  $\text{NaNO}_3$ .
- (ii) Many candidates differentiated on the basis of smell and did not fulfil the requirement of chemical test. Some suggested blue litmus turning red for  $\text{HCl}$  and the other way round for  $\text{H}_2\text{S}$ . Few of the candidates did not use the same reagent for both gases.
- (iii) Many candidates brought out the difference on the basis of bonds present i.e. double bond and single bonds and did not give a chemical test to distinguish. Some other candidates specified the kind of reactions shown by both i.e. addition and substitution reactions respectively or the type of compounds as being unsaturated and saturated.
- (iv) Candidates used  $\text{NaOH}$  as the distinguishing reagent and stated the colour of the ppt with each solution but failed to specify the solubilities of the ppt in the reagent.  
Many candidates were not aware that  $\text{Ca}(\text{NO}_3)_2$  solution does not form ppt with  $\text{NH}_4\text{OH}$ .
- (v) Many candidates did not associate the word solution / paper / acidified with  $\text{K}_2\text{Cr}_2\text{O}_7$  or  $\text{KMnO}_4$ .  
Some others used lime water test which was answered by both gases.
- (f) (ii) Chlorine and lithium were the occasional wrong answers.
- (iii) Most attempted correctly. Few candidates chose Ammonia as the correct options.
- (iv) Most candidates answered correctly.
- (v) Some of the candidates wrote acetic acid.
- (vi) Most of the candidates answered correctly.
- (vii) Oxide chosen at random.
- (viii) Some incorrectly answered this part as dehydrating agent.
- (ix) Generally correctly answered. Few candidates incorrectly answered as Ammonium nitrate.
- (x) Most candidates answered this part correctly. Occasional errors showed a lack of understanding of the properties of ionic compounds.
- (xi) There was confusion between the components of brass and bronze.
- (g) (i) Candidates failed to calculate the gram molecular weight correctly.
- (ii) Knowledge of gram molecular volume was missing.
- (iii) Most candidates answered this part correctly. Some suggested  $6.023 \times 10^{23}$  molecules of  $\text{N}_2$  or multiplied  $\text{N}$  by  $6.023 \times 10^{23}$  molecules.

#### Suggestions for teachers

- Students must be clear about the basic concepts so that understanding relation between mole, molar mass and molar volume is possible.
- Students must be given adequate practice in application of Avogadro's law after explaining it and pointing out the variations from Gay Lussac's law

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**L Darling-Hammond**



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