Chapter 24 - Amines and Heterocycles

Chapter Outline

- Facts about amines (Section 24.1-24.5).
 - A. Naming amines (Section 24.1).
 - Amines are classified as primary (RNH₂), secondary (R₂NH), tertiary (R₂N) or quaternary ammonium salts (R₄N⁺).
 - 2. Primary amines are named in several ways:
 - For simple amines, the suffix -amine is added to the name of the alkyl substituent.
 - The suffix, -arraine can replace the final -e of the parent compound.
 - For more complicated amines, the -NH₂ group is an amino substituent on the parent molecule.
 - Secondary and tertiary amines:
 - Symmetrical amines are named by using the prefixes all- and tri- before the name of the alkyl group.
 - b. Unsymmetrical arrines are named as N-substituted primary arrines.
 - i. The largest group is the parent,
 - The simplest arylamine is aniline.
 - Heterocyclic amines (nitrogen is part of a ring) have specific parent names.
 The nitrogens receive the lowest possible numbers.
 - B. Structure and properties of amines (Section 24.2).
 - 1. The three amine bonds and the lone pair occupy the corners of a tetrahedron.
 - 2. An amine with three different substituents is chiral.
 - The two amine enantiomers interconvert by pyramidal inversion.
 - This process is rapid at room temperature.
 - Amines with fewer than 5 carbons are water-soluble and form hydrogen bonds.
 - 4. Amines have higher boiling points than alkanes of similar molecular weight.
 - Amines smell nasty!
 - C. Basicity of amines (Sections 24.3–24.5).
 - The lone pair of electrons makes amines both nucleophilic and basic (Section 24.3).
 - The basicity constant K₈ is the measure of the equilibrium of an amine with water.
 - The larger the value of K_b (smaller pK_b), the stronger the base.
 - More often, K_a is used to describe amine basicity.
 - K_k is the dissociation constant of the conjugate acid of an amine.
 - b. $pK_a + pK_b = 14$ (for aqueous media).
 - The smaller the value of K_n (larger pK_n), the stronger the base.
 - 4. Base strength.
 - Primary, secondary, and tertiary alkylamines have similar basicities.
 - b. Arylamines and beterocyclic amines are less basic than alkylamines.
 - The xp² electrons of the pyridine lone pair are less available for bonding.
 - The pyrrole lone pair electrons are part of the aromatic ring a system.
 - c. Amides are nonbasic.
 - d. Amine busicity can be used as a means of separating amines from a mixture.
 i. An amine can be converted to its salt, extracted from an organic solution with water, neutralized, and re-extracted with an organic solvent.
 - e. Some amines are very weak acids.
 - LDA is formed from disopropylamine and acts as a strong base.

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