

### balance nuclear reaction equations

aluminum-26 undergoes positron emission  
potassium-40 undergoes beta decay (electron)  
uranium-238 undergoes alpha decay



### electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



### NOMENCLATURE [naming compounds, writing formulas]

Write the formula for each of the following compounds:

- a. Sulfur hexafluoride  $\text{SF}_6$
- b. Lithium nitride  $\text{Li}_3\text{N}$
- c. Chromium (III) carbonate  $\text{Cr}_2(\text{CO}_3)_3$
- d. Tin (II) chloride  $\text{SnCl}_2$
- e. Ammonium acetate  $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$  (or  $\text{CH}_3\text{COONH}_4$ )
- f. Mercury (I) chloride  $\text{Hg}_2\text{Cl}_2$
- g. Potassium bromate  $\text{KBrO}_3$
- h. Hydrosulfuric acid  $\text{H}_2\text{S(aq)}$
- i. Chloric acid  $\text{HClO}_3$
- j. Sulfurous acid  $\text{H}_2\text{SO}_3$

Name each of the following compounds.

- a.  $\text{CuSO}_4$  Copper (II) sulfate
- b.  $\text{AlF}_3$  aluminum fluoride
- c.  $\text{HI}$  hydrogen iodide
- d.  $\text{NO}$  nitrogen monoxide
- e.  $\text{H}_2\text{Se(aq)}$  hydroselenic acid
- f.  $\text{HNO}_3$  nitric acid
- g.  $\text{HNO}_2$  nitrous acid
- h.  $\text{NaHSO}_4$  sodium hydrogen sulfate

### Half-life

Tritium ( $\text{H-3}$ ) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \xrightarrow{1} 20 \xrightarrow{2} 10 \xrightarrow{3} 5 \xrightarrow{4} 2.5 \xrightarrow{5} 1.25$$

**61.5 yr**

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \xrightarrow{1} 50 \xrightarrow{2} 25 \xrightarrow{3} \boxed{12.5 \text{ mg}}$$

### Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ \hline & + & \\ & 63.62 & \rightarrow 64 \end{array}$$

# Measurements And Calculations Chemistry Study Guide Answers

**Peter Scott, Rachael Enriquez Porter**



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## **Table of Contents Measurements And Calculations Chemistry Study Guide Answers**

1. Understanding the eBook Measurements And Calculations Chemistry Study Guide Answers
  - The Rise of Digital Reading Measurements And Calculations Chemistry Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Measurements And Calculations Chemistry Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Measurements And Calculations Chemistry Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Measurements And Calculations Chemistry Study Guide Answers
  - Personalized Recommendations
  - Measurements And Calculations Chemistry Study Guide Answers User Reviews and Ratings
  - Measurements And Calculations Chemistry Study Guide Answers and Bestseller Lists

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  - Measurements And Calculations Chemistry Study Guide Answers Enhanced eBook Features
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  - Highlighting and Note-Taking Measurements And Calculations Chemistry Study Guide Answers
  - Interactive Elements Measurements And Calculations Chemistry Study Guide Answers
8. Staying Engaged with Measurements And Calculations Chemistry Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Measurements And Calculations Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Measurements And Calculations Chemistry Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Measurements And Calculations Chemistry Study Guide Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Measurements And Calculations Chemistry Study Guide Answers
  - Setting Reading Goals Measurements And Calculations Chemistry Study Guide Answers
  - Carving Out Dedicated Reading Time
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  - Fact-Checking eBook Content of Measurements And Calculations Chemistry Study Guide Answers
  - Distinguishing Credible Sources
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- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
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