Honors Chemistry Lab 7   Electron Configurations of atoms and ions
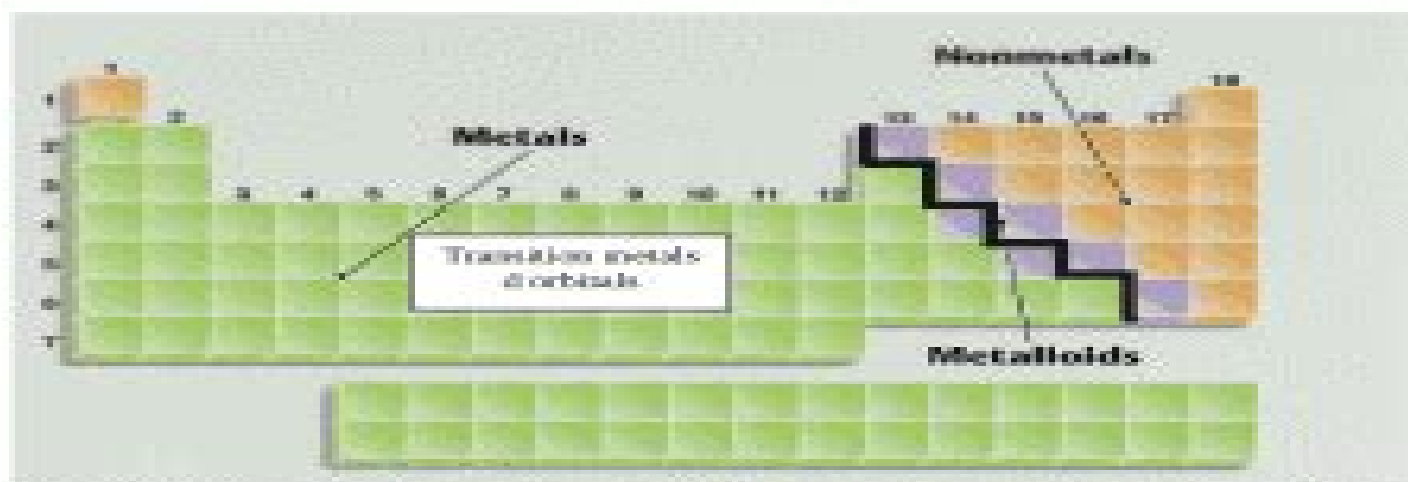
Objectives
- Identify the metals and nonmetals on a periodic table
- Observe colors of various chemical solutions
- Write electron configurations for various metals, nonmetals, cations, and anions
- Relate the presence of color in an ionic solution as a characteristic of electron configuration

Introduction:

Nonmetals

Metals

Transition metals
d orbitals

Metalloids

Notice the location of Metals and Nonmetals. Metals tend to have relatively **low ionization energies**. Thus, they generally lose electrons readily. Loss of electrons by a metal atom produces a positively charged ion called a **cation**. The electrical charge of a metal cation depends on the number of electrons the metal loses. When they form cations, they usually lose their valence electrons in the s and p subleveds. Upon producing a cation, these metals end up with positive charges group 1 = 1⁺; group 2⁺; group 5= 3⁺; group 6=2⁺; group 7=3⁺. Upon losing all its valence electrons, the metal cation generally has the stable outer electron configuration of a noble gas. For example when potassium becomes K⁺ its electron configuration is the same as argon, with 18 electrons.

Nonmetals have relatively **high ionization energies**, so they tend not to lose electrons. Instead they gain electrons to produce negative **anions**. They generally gain just enough electrons to attain the stable electron configuration of a noble gas, or eight outer energy level electrons in their highest occupied s and p orbitals. For example bromine will gain one more electron (Br⁻) to have 36 electrons, just like Kr, a noble gas.

In general, to write the electron configuration of any cation or anion, write the electron configuration for the neutral atom first, then add or subtract the required number of electrons. Feel free to write place holders such as 4s⁰ showing that there are 0 electrons in the 4s, or just eliminate the 4s from the electron configuration.

Transition metals also lose electrons to produce cations. Unlike most representative metals, the number of electrons that can be lost by transitions metals can vary. Iron for example, can for the iron 3⁺ ion and the 2⁺ ion, by losing three or two electrons. In general, **transition metals lose their outer s orbital electrons before**

Lab Electron Configurations Of Atoms And Ions

Richard Bailey

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