

Properties of Acids and Bases

(Adapted from Flinn Scientific Acid - Base Test Kit I - #AP4567)

Introduction

Battery acid, stomach acid, acid rain - just a few acids in our everyday life! What does it mean when something is acidic? How do acids differ from bases? In this lab, the physical and chemical properties of a variety of acids and bases will be examined.

Concepts

- Acids vs. Bases
- Indicators
- pH

Background

Acids and bases make up two groups of substances that can be categorized by their physical and chemical properties. Let's take a look at the distinguishing properties:

1. **Taste:** Acids, such as lemons or oranges, taste sour. Bases, such as soap, taste bitter. *Note:* Taste should never be used to identify a lab chemical or unknown substance. This lab does not include a test using taste.
2. **Feel:** Acid solutions do not feel much different than water; however, they sting if they contact broken skin. Base solutions have a slippery feeling. *Note:* Lab chemicals or unknown substances should never be touched with the bare skin. This lab does not include a test using feel.
3. **Indicators:** An indicator is a chemical compound, either on a test paper or in a solution, that changes color depending on the acidity or basicity of a solution and, thus, is used to test for the presence of acids or bases. There are many different kinds of indicators. Litmus paper and phenolphthalein indicator solution are two of the most common.
 - Blue litmus paper turns red when dipped in an acid solution of pH 1-6.
 - Red litmus paper turns blue when dipped in a basic solution of pH 8-14.
 - In a neutral solution with pH 7, each paper retains its original color.
 - Phenolphthalein indicator solution turns a pink color in a basic solution but remains colorless (sometimes cloudy) in an acidic or neutral solution.
4. **Reaction with carbonates:** While most bases do not react with carbonates, acids react with carbonates to release carbon dioxide gas. For example:
$$\text{BaCO}_3 + 2\text{HCl} \rightarrow \text{BaCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$$
5. **Reaction with metals:** While most bases do not react with metals, acids react with active metals to release hydrogen gas. For example:
$$\text{Fe} + 2\text{HCl} \rightarrow \text{FeCl}_2 + \text{H}_2$$

Lab Properties Of Acids And Bases Answers

KJ Lindholm-Leary



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Table of Contents Lab Properties Of Acids And Bases Answers

1. Understanding the eBook Lab Properties Of Acids And Bases Answers
 - The Rise of Digital Reading Lab Properties Of Acids And Bases Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Properties Of Acids And Bases Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Properties Of Acids And Bases Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Properties Of Acids And Bases Answers
 - Personalized Recommendations
 - Lab Properties Of Acids And Bases Answers User Reviews and Ratings
 - Lab Properties Of Acids And Bases Answers and Bestseller Lists
5. Accessing Lab Properties Of Acids And Bases Answers Free and Paid eBooks

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 - Lab Properties Of Acids And Bases Answers Budget-Friendly Options
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 - Lab Properties Of Acids And Bases Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Lab Properties Of Acids And Bases Answers
 - Highlighting and Note-Taking Lab Properties Of Acids And Bases Answers
 - Interactive Elements Lab Properties Of Acids And Bases Answers
 8. Staying Engaged with Lab Properties Of Acids And Bases Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Properties Of Acids And Bases Answers
 9. Balancing eBooks and Physical Books Lab Properties Of Acids And Bases Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Properties Of Acids And Bases Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab Properties Of Acids And Bases Answers
 - Setting Reading Goals Lab Properties Of Acids And Bases Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab Properties Of Acids And Bases Answers
 - Fact-Checking eBook Content of Lab Properties Of Acids And Bases Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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